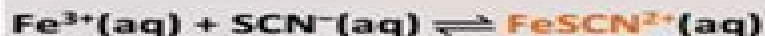


IA lab work on spectrophotometric determination of K_c of the Iron(3)Thiocyanate system



$$K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$$

Since I will not change the temperature in my experiment, the equilibrium constant will stay the same, and therefore I can measure the light absorbance and use the Lambert Beer-Law to find the initial concentration of each species in the system and then calculate the K_c .

$$X(\text{initial concentration of Iron3 and Thiocyanate}) = A(\text{absorbance of Iron(3)Thiocyanate solution}) / \epsilon \cdot l$$

Equipment:

- 4 test tubes with 4 corks
- test tube rack
- 10 ml pipette with pump thing

Solutions:

- 20 ml of Iron (3) solution
- 20 ml of Thiocyanate solution
- 20 ml of water H_2O

Samples	The amount of Thiocyanate solution SCN^{-}	The amount of water H_2O	Iron 3 solution Fe^{3+}
1	2 ml	8 ml	5 ml
2	4 ml	6 ml	5 ml
3	5 ml	5 ml	5 ml
4	9 ml	1 ml	5 ml

Kc Iron Thiocyanate Lab Report

Lei Huang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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 - The Rise of Digital Reading Kc Iron Thiocyanate Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Kc Iron Thiocyanate Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kc Iron Thiocyanate Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kc Iron Thiocyanate Lab Report

- Personalized Recommendations
- Kc Iron Thiocyanate Lab Report User Reviews and Ratings
- Kc Iron Thiocyanate Lab Report and Bestseller Lists
- 5. Accessing Kc Iron Thiocyanate Lab Report Free and Paid eBooks
 - Kc Iron Thiocyanate Lab Report Public Domain eBooks
 - Kc Iron Thiocyanate Lab Report eBook Subscription Services
 - Kc Iron Thiocyanate Lab Report Budget-Friendly Options
- 6. Navigating Kc Iron Thiocyanate Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Kc Iron Thiocyanate Lab Report Compatibility with Devices
 - Kc Iron Thiocyanate Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kc Iron Thiocyanate Lab Report
 - Highlighting and Note-Taking Kc Iron Thiocyanate Lab Report
 - Interactive Elements Kc Iron Thiocyanate Lab Report
- 8. Staying Engaged with Kc Iron Thiocyanate Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kc Iron Thiocyanate Lab Report
- 9. Balancing eBooks and Physical Books Kc Iron Thiocyanate Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kc Iron Thiocyanate Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Kc Iron Thiocyanate Lab Report
 - Setting Reading Goals Kc Iron Thiocyanate Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Kc Iron Thiocyanate Lab Report

- Fact-Checking eBook Content of Kc Iron Thiocyanate Lab Report
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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