

Objective type

- Q.1 Which of the following is not a general characteristic of equilibria involving physical processes?
- (1) Equilibrium is possible only in a closed system at a given temperature.
 - (2) All measurable properties of the system remain constant.
 - (3) All the physical processes stop at equilibrium.
 - (4) The opposing processes occur at the same rate and there is dynamic but stable condition.
- Q.2 Which of the following statements is incorrect?
- (1) In equilibrium mixture of ice and water kept in perfectly insulated flask mass of ice and water does not change with time.
 - (2) The intensity of red colour increases when oxalic acid is added to a solution containing iron (III) nitrate and potassium thiocyanate.
 - (3) On addition of catalyst the equilibrium constant value is not affected.
 - (4) Equilibrium constant for a reaction with negative ΔH value decreases as the temperature increases.
- Q.3 When hydrochloric acid is added to cobalt nitrate solution at room temperature, the following reaction takes place and the reaction mixture becomes blue. On cooling the mixture it becomes pink. On the basis of this information mark the correct answer.
- $$\underset{\text{(pink)}}{[\text{Co}(\text{H}_2\text{O})_6]^{3+}}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \underset{\text{(blue)}}{[\text{CoCl}_4]^{2-}}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$$
- (1) $\Delta H > 0$ for the reaction
 - (2) $\Delta H < 0$ for the reaction
 - (3) $\Delta H = 0$ for the reaction
 - (4) The sign of ΔH cannot be predicted on the basis of this information.
- Q.4 What will be the correct order of vapour pressure of water, acetone and ether at 30°C . Given that among these compounds, water has maximum boiling point and ether has minimum boiling point?
- (1) Water < ether < acetone
 - (2) Water < acetone < ether
 - (3) Ether < acetone < water
 - (4) Acetone < ether < water
- Q.5 In which of the following reactions, the equilibrium remains unaffected on addition of small amount of argon at constant volume?
- (1) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 - (2) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
 - (3) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
 - (4) The equilibrium will remain unaffected in all the three cases

Assertion & Reason

- Q.6 Assertion (A): In the dissociation of PCl_5 at constant pressure and temperature addition of helium at equilibrium increases the dissociation of PCl_5 .
- Reason (R) : Helium removes Cl_2 from the field of action.
- (1) Both A and R are true and R is correct explanation of A.
 - (2) Both A and R are true but R is not correct explanation of A.
 - (3) A is true but R is false.
 - (4) Both A and R are false.

Mcqs Of Environmental Chemistry

Michal Rosen-Zvi



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Understanding our Environment R M Harrison, 2007-10-31 This 2nd edition of

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Environmental Chemistry John Moore, 2012-12-02 Environmental Chemistry concerns with the broad interpretation on what environmental chemistry is and discusses chemistry in relation to environmental topics The book is divided into seven parts Part I discusses the origins of different elements and interstellar molecules the development of the earth and the chemical evolution of life Part II talks about energy and its theoretical treatment the origin development and problems related to fossil fuels and the developing energy sources including storage distribution and conservation Part III discusses the air the structure and properties of the atmosphere and air pollution in relation to different industries and transportation Mineral resources and solid wastes are tackled in Part IV and the principles and treatment of water are explained in Part V Part VI discusses the sustenance of life amino acids and the control of toxins and Part VII studies the relationship of science ethics and ecology The text is good for those in the field of chemistry and wish to understand the importance of their field to the environment and for environmentalists and ecologists who want to know the relationship of chemistry with their studies

Environmental Chemistry , **Annual Report - University of the Witwatersrand** University of the Witwatersrand, **Key Concepts in Environmental Chemistry** Grady Hanrahan, 2011-08-02 Key Concepts in Environmental Chemistry provides a modern and concise introduction to environmental chemistry principles and the dynamic nature of environmental systems It offers an intense one semester examination of selected concepts encountered in this field of study and provides integrated tools in explaining complex chemical problems of environmental importance Principles typically covered in more comprehensive textbooks are well integrated into general chapter topics and application areas The goal of this textbook is to provide students with a valuable resource for learning the basic concepts of environmental chemistry from an easy to follow condensed application and inquiry based perspective Additional statistical sampling modeling and data analysis concepts and exercises will be introduced for greater understanding of the underlying processes of complex environmental systems and fundamental chemical principles Each chapter will have problem oriented exercises with examples throughout the body of the chapter that stress the important concepts covered and research applications case studies from experts in the field Research

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Environmental Chemistry John Wright,2004-03 Offers an accessible introduction to chemical principles and concepts and makes the subject accessible to those with little or no previous knowledge of chemistry It is highly illustrated with global case studies figures and tables

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