#### PART A -CHEMISTRY

- Total An unknown alcohol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism:
  - tertiary alcohol by S<sub>N</sub>1.

(2) secondary alcohol by S<sub>v</sub>2

(3) tertiary alcohol by S<sub>x</sub>2

(4) secondary alcohol by S<sub>n</sub>1

Seed.

Reaction proceeds through carbocation formation as 3° carbocation is highly stable, hence reaction proceeds through S<sub>N</sub>1 with 3" alcohol.

20 The first ionization potential of Na is 5.1 eV. The value of electron gain enthalpy of Na\* will be:

(0)	- 5.	LeV
-----	------	-----

(3) + 2.55 eV

Seed.

Na Na Na Na Na Na + e , here the backward reaction releases same amount of energy and known as

Electron gain outhalpy,

Stability of the species Li<sub>2</sub>, Li<sub>2</sub> and Li<sub>3</sub> increases in the order of: 200

(1) 
$$\text{Li}_{2}^{2} < \text{Li}_{2}^{2} < \text{Li}_{2}$$

(3) 
$$Li_2 < Li_2 < Li_2$$

(4) 
$$Li_2 < Li_2^2 < Li_2^2$$

Sect.

$$\text{Li}_2(6) = \sigma \text{Is}^2 \ \hat{\sigma} \text{Is}^2 \ \sigma 2\text{s}^2$$

B.O. = 
$$\frac{4-2}{2} = 1$$

$$Li_{5}^{2}(5) = \sigma i s^{2} \dot{\sigma} i s^{2} \sigma 2 s^{1}$$

B. O. 
$$=\frac{3-2}{2}=0.5$$

$$\text{Li}_{2}^{2}(7) = \sigma \text{Is}^{2} \hat{\sigma} \text{Is}^{2} \sigma 2\text{s}^{2} \hat{\sigma} 2\text{s}^{4}$$

B.O. 
$$=\frac{4-3}{2}=0.5$$

Li," is more stable than Li, because Li, has more numbers of antibonding electrons.

- 400 The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2(M)HCl will be:
  - CES 1,000 MG (30: 0.975 M

Sect.

Seed.  $M_1V_1 + M_2V_2 = MV$  $M = \frac{M_1V_1 + M_2V_2}{M_1} = \frac{0.5 \times 750 + 2 \times 250}{M_2}$ 

5 Which of the following is the wrong statement?

(1) O<sub>1</sub> molecule is bent

- (2) Ozone is violet-black in solid state (4) ONCI and ONO' are not isoelectronic
- (3) Ozone is diamagnetic gas (All the options are correct statements)

(1) Correct, as O O is bent

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