

PHYSICAL KINETIC MOLECULAR THEORY

Learning Objectives

- 1) Identify theoretical differences in how gases participate in the kinetic theory and how molecules in liquids and solids participate in the kinetic molecular theory
- 2) Identify how temperature affects molecular motion
- 3) Explain how kinetic molecular theory is applied to the behavior of everyday substances

Module 1: Representation of Gases in Different Phases

The following figure illustrates the representative particle for the three phases of matter: solid, liquid, and gas.



Figure 1: Representative particles for the three phases of matter.

Understanding the differences in how gases, liquids, and solids participate in the kinetic molecular theory is essential for understanding the behavior of matter in different phases.

Key Concepts

1. Matter can be in three phases: solid, liquid, and gas. Each phase has different properties, and the particles in each phase behave differently. The kinetic molecular theory explains the behavior of matter in different phases.

Table 1: Kinetic Molecular Theory in Different Phases

	SOLID	LIQUID	GAS
Particle Arrangement	Particles are packed closely together in a fixed, ordered arrangement. The particles are in contact with each other.	Particles are packed closely together in a disordered arrangement. The particles are in contact with each other.	Particles are widely spaced and move freely in all directions. The particles are not in contact with each other.
Particle Motion	Particles vibrate in place. They have low kinetic energy and move slowly.	Particles move more freely than in the solid phase. They have moderate kinetic energy and move faster.	Particles move very freely. They have high kinetic energy and move very fast.

Lw 1 Pogil Kinetic Molecular Theory Key

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