

i.MX Linux® User's Guide

1 Overview

This document describes how to build and install the i.MX Linux® OS BSP, where BSP stands for Board Support Package, on the i.MX platform. It also covers special i.MX features and how to use them.

The document also provides the steps to run the i.MX platform, including board DIP switch settings, and instructions on configuring and using the U-Boot bootloader.

The later chapters describe how to use some i.MX special features when running the Linux OS kernel.

Features covered in this guide may be specific to particular boards or SOC's. For the capabilities of a particular board or SOC, see the *i.MX Linux® Release Notes* (IMXLXRN).

1.1 Audience

This document is intended for software, hardware, and system engineers who are planning to use the product, and for anyone who wants to know more about the product.

1.2 Conventions

This document uses the following conventions:

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Mx 19 User Guide

Marcel A. Müller



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web apr 15 2022 creating a single wiring diagram for a central lock system involves the use of several components depending on the specifics of the system you may need to include relays electric motors switches transponders and door locks

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increasingly sophisticated and are now commonly used by automotive repair professionals and enthusiasts alike

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web installing wires make sure that all actuators are installed and connected before connecting red and black wire to power supply 12v and ground during connecting bunch of wires to front door actuators 5 wires make sure they are connected correctly taking into account colors see figure 5 front actuator bunch white brown black controls

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web check the functionality of the central locking system by closing and re opening the doors with your car key lock the vehicle again current supply red cable at constant 12v black cable at ground connect to permanent 12v vehicle fused 12v constant attach to battery minus to the vehicle chassis ground

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant k is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined

with eqn 2 to determine Δh and Δs from the temperature dependence of k

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web sep 18 2020 in this experiment we will titrate a saturated calcium hydroxide solution with a standardized hcl solution we will determine the hydroxide ion concentration for the saturated calcium hydroxide

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared Ca(OH)_2 solution with a standard hydrochloric acid solution

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the ksp of a slightly soluble basic salt and how solubility changes with temperature ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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molarity 0.050 M volume of filtered Ca(OH)_2 solution 50.05 mL 3.71375 equivalence point volume mL in saturated Ca(OH)_2 in saturated Ca(OH)_2 K_{sp} of Ca(OH)_2 average K_{sp} of Ca(OH)_2

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be K_{sp} Ca(OH)_2 from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant K_{sp} chemistry libretexts - Feb 01 2023

web Jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determining the K_{sp} of calcium hydroxide the purpose of this was to titrate a saturated Ca(OH)_2 solution with a standard HCl solution and find the K_{sp} or the OH^- concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the K_{sp} of calcium hydroxide $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ for this experiment the titration of calcium hydroxide with hydrochloric acid was investigated determining the K_{sp} of calcium hydroxide $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ for this experiment the titration of calcium

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web question chem 182 experiment 8 determining the K_{sp} of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form solution of calcium ionic solid that is sparingly soluble in water a saturated aqueous as shown below

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest K_{sp} value to the theoretical K_{sp} for Ca(OH)_2 was the determination through pH method 2 introduction calcium hydroxide is a strong base soluble in water

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web plasma membrane homework symbiosis practice worksheet benzoin condensation grade A determining the K_{sp} of calcium hydroxide after reading the lab complete items a b c and d title purpose chemicals and equipment and summary of procedure as described on page 10 of exp 1 on an 8.5 x 11 sheet of paper

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web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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web calculate the ksp for calcium hydroxide Ca^{2+} OH^- 2.0 x 10⁻² ksp 5 x 10⁻⁴ find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5 x 10⁻⁶

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