

CHM 2210 Lab: Melting Point

Introduction

The melting point of a solid is the temperature at which the solid melts (changes to a liquid) at 1 atm of pressure. Melting point is determined by slow heating of a small solid sample placed in a capillary tube. The melting point should be reported as a melting point range: from the temperature at which the first droplet of liquid is observed to the temperature at which the sample becomes a clear liquid, for example: 96.7 – 98.9°C.

A pure solid (pure enough for most purposes) usually has a sharp melting point which means that it melts within a range of 2°C or less.

An impure solid melts at a broader range (3°C or more) and at depressed temperature.

Using melting point to find the identity of a solid is not sufficient because there are too many substances with the same melting points.

However, you can use a mixed melting point as a tool for identification of a solid.

If you mix two different solids with the same melting point, the melting point of the mixture will be broad and depressed. The melting point of a solid and a mixture of this solid with the authentic solid should be the same.

Example:

Measured/Observed melting point of an unknown solid X: 128.0 – 129.5°C

Melting point of known solid A: 130°C

Melting point of known solid B: 158°C

Measured melting point of mixture X + A: 127.5 – 129.0°C

Measured melting point of mixture X + B: 133.5 – 142.0°C

Identity of unknown X: solid A

In the future you will measure melting points of solids mostly to determine their purity.

- **Preparation of Sample**

A solid sample should be dried and crushed with a metal spatula or bottom of a test tube.

If a solid is used for mixed melting point, the 1:1 mixture should be crushed together into a homogeneous mixture.

The pulverized solid should be mounded up and a small (1 – 2 mm) plug of sample should be pushed into the open end of the capillary tube. The sample should be knocked down to the sealed end of the capillary tube by tapping its sealed end on the surface or dropping it into a long glass tubing.

The sample should be 1-2 mm long and packed firmly.

If the sample is too large, the measured melting point may be too high.

If the sample is loosely packed, it may be difficult to determine the melting point range because of the air present.

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Daniela Niemeyer

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