Chapter test with answers

Chapter 9 Reaction rates

Time permitted: 30 minutes

	Section	Number of questions	Marks available
А	Multiple choice	15	15
В	Short answer	5	15
	Total		30

Seales

A+	29-30	A	26-28	В	23-25	C	19-22	D	15-18	E	9-14	uG	0-8
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Section A Multiple choice (15 marks)

Section A consists of 15 questions, each worth one mark. Each question has only one correct answer. Circle the correct answer. Attempt all questions. Marks will not be deducted for incorrect answers. You are advised to spend no more than 15 minutes on this section.

- 1 Why does reaction rate increase as temperature increases?
 - Particles move faster and this increases the probability of collision.
 - B There are more particles at higher temperatures so more of them collide.
 - C The activation energy is lowered.
 - D Particles require a set temperature in which to begin colliding.

For questions 2 and 3 refer to the chemical reaction below.

$$CaCO_3(s) + 2H^*(aq) \rightarrow Ca^{2*}(aq) + CO_2(g) + H_2O(l)$$

- 2 Which does not occur during the reaction?
 - The concentration of H₂O increases.
 - B The mass of CaCO₃ decreases.
 - C The volume of CO2 increases.
 - D The concentration of Ca2+ increases.
- 3 The rate of the reaction could not be determined by:
 - A recording the change in mass of the CaCO3 over time.
 - B recording the mass of CO₂ lost over time.
 - measuring the final mass of CaCO₃.
 - D collecting and recording the volume of CO2 gas produced over time.

Lesson 36 Reaction Rates Answers Key

Leonard Glick

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