

Chapter test with answers

Chapter 9 Reaction rates

Time permitted: 30 minutes

	Section	Number of questions	Marks available
A	Multiple choice	15	15
B	Short answer	5	15
	Total		30

Scale:

A+	29–30	A	26–28	B	23–25	C	19–22	D	15–18	E	9–14	UG	0–8
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Section A Multiple choice (15 marks)

Section A consists of 15 questions, each worth one mark. Each question has only one correct answer. Circle the correct answer. Attempt all questions. Marks will not be deducted for incorrect answers. You are advised to spend no more than 15 minutes on this section.

- 1 Why does reaction rate increase as temperature increases?
- ☒ **A** Particles move faster and this increases the probability of collision.
 - ☐ **B** There are more particles at higher temperatures so more of them collide.
 - ☐ **C** The activation energy is lowered.
 - ☐ **D** Particles require a set temperature in which to begin colliding.

For questions 2 and 3 refer to the chemical reaction below.



- 2 Which does not occur during the reaction?
- ☒ **A** The concentration of H_2O increases.
 - ☐ **B** The mass of CaCO_3 decreases.
 - ☐ **C** The volume of CO_2 increases.
 - ☐ **D** The concentration of Ca^{2+} increases.
- 3 The rate of the reaction could not be determined by:
- ☐ **A** recording the change in mass of the CaCO_3 over time.
 - ☐ **B** recording the mass of CO_2 lost over time.
 - ☒ **C** measuring the final mass of CaCO_3 .
 - ☐ **D** collecting and recording the volume of CO_2 gas produced over time.

Lesson 36 Reaction Rates Answers

S. Scott Sutton

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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