

Calculations in Analytical Chemistry

The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is, $N_A = 6.022 \times 10^{23}$.

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

| Element | Molar mass per element | Number of atoms | Mass |
|------------|------------------------|-----------------|-----------|
| C | 12 g/mol | 6 | 72 g/mol |
| H | 1 g/mol | 12 | 12 g/mol |
| O | 16 g/mol | 6 | 96 g/mol |
| Molar Mass | | | 180 g/mol |

Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

1. Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 g_{HBz} \times \frac{1 mol_{HBz}}{122.1 g_{HBz}} = 0.0164 mol_{HBz}$$

To obtain the number of millimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 g_{HBz} \times \frac{1 mmol_{HBz}}{0.1221 g_{HBz}} = 16.4 mmol_{HBz}$$

2. What is the mass in grams of Na^+ (22.99 g/mol) in 25.0 g of Na_2SO_4 (142.0 g/mol)?

$$mass_{Na^+} = 25.0 g_{Na_2SO_4} \times \frac{1 mol_{Na_2SO_4}}{142.0 g_{Na_2SO_4}} \times \frac{2 mol_{Na^+}}{1 mol_{Na_2SO_4}} \times \frac{22.99 g_{Na^+}}{1 mol_{Na^+}}$$

$$mass_{Na^+} = 8.10 g_{Na^+}$$

Manual Calculations Of Food Analytical Chemistry

B. Pam Ismail, S. Suzanne Nielsen



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A Manual of Analytical Chemistry, Qualitative and Quantitative--inorganic and Organic John Muter, 1887 *Handbook of Food Analysis - Two Volume Set* Leo M.L. Nollet, Fidel Toldra, 2015-06-10 Updated to reflect changes in the industry during the last ten years The Handbook of Food Analysis Third Edition covers the new analysis systems optimization of existing techniques and automation and miniaturization methods Under the editorial guidance of food science pioneer Leo M L Nollet and new editor Fidel Toldra the chapters take an in *Nielsen's Food Analysis* B. Pam Ismail, S. Suzanne Nielsen, 2024-06-24 This sixth edition provides information on techniques needed to analyze foods for chemical and physical properties The book is ideal for undergraduate courses in food analysis and it is also an invaluable reference for professionals in the food industry General information chapters on regulations labeling sampling and data handling provide background information for chapters on specific methods to determine chemical composition and characteristics physical properties and constituents of concern Methods of analysis cover information on the basic principles advantages limitations and applications The information on food analysis applications has been expanded in a number of chapters that cover basic analytical techniques Instructors who adopt the textbook can contact B Ismail for access to a website with related teaching materials Analytical Chemistry Frederick Pearson Treadwell, 1904 Technical Methods of Ore Analysis Albert Howard Low, 1911

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