

6. When one gram acid of  $\text{KMnO}_4$  reacts with  $\text{HCl}$ , the volume of solution liberated at NTP will be
- (a) 11.2 litres (b) 22.4 litres  
(c) 44.8 litres (d) 56.8 litres
7. When  $\text{MnO}_4^-$  ions reacts with  $\text{Fe}^{2+}$  ions in acid solution the  $\text{Fe}^{3+}$  is liberated. The equivalent weight of  $\text{KMnO}_4$  in this reaction is
- (a) 500 (b) 1500  
(c) 300 (d) 2000
8. 54 g. of hydrogen peroxide is present in 1000 ml. of solution. This solution is called
- (a) 10 vol solution (b) 20 vol solution  
(c) 30 vol solution (d) 50 vol solution
9. The number of moles of  $\text{KMnO}_4$  that will be needed to react completely with one mole of ferrous oxalate in acidic solution is
- (a) 100 (b) 200  
(c) 400 (d) 2
10. To prepare a solution that is 0.05 M  $\text{KCl}$  starting with 100 mL of 0.40 M  $\text{KCl}$
- (a) add 0.30 g  $\text{KCl}$   
(b) add 50 mL of water  
(c) add 0.30 mol  $\text{KCl}$   
(d) evaporate 10 mL water
11. In acid solution solution,  $\text{Br}_2$  disproportionates to  $\text{Br}^-$  and  $\text{BrO}_3^-$
- $$3\text{Br}_2 + 6\text{H}_2\text{O} \rightarrow 2\text{Br}^- + \text{BrO}_3^- + 6\text{H}^+$$
- Atomic equivalent weight of  $\text{Br}_2$  is (molecular weight = 160)
- (a)  $\frac{16}{3}$  (b)  $\frac{16}{2}$   
(c)  $\frac{160}{3}$  (d)  $\frac{160}{2}$
12.  $(\text{Fe}^{2+})$  in a solution prepared by adding 20.00 mL of 0.10 M  $\text{FeSO}_4$  with 70 mL of 0.10 M  $\text{Na}_2\text{SO}_4$  is
- (a) 0.020 M (b) 0.140 M  
(c) 0.010 M (d) 0.200 M
13. When 80 mL of 0.100 M  $\text{HCl}$  is mixed with 120 mL of 0.10 M  $\text{KOH}$ , the resultant solution is the same as a solution of
- (a) 0.10 M  $\text{HCl}$  and 0.02 M  $\text{HCl}$   
(b) 0.08 M  $\text{HCl}$   
(c) 0.08 M  $\text{HCl}$  and 0.01 M  $\text{KOH}$   
(d) 0.08 M  $\text{HCl}$  and 0.01 M  $\text{HCl}$
14. 50 mL of  $\text{H}_2\text{O}_2$  solution contains strength = x% required 10 mL of  $\text{KMnO}_4$  solution in acidic medium. Hence x is
- (a) 0.20 (b) 0.4  
(c) 0.5 (d) 06.8
15. Molarity of 1.0 M  $\text{H}_2\text{SO}_4$  (d = 1.8 mL/cm<sup>3</sup>) is
- (a) 36 mol kg<sup>-1</sup> (b) 300 mol kg<sup>-1</sup>

# Multiple Choice Questions On Inorganic Chemistry

**Jonathan Orsay**



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