PART A -CHEMISTRY

- An unknown alcohol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism:
 - (I) tertiary alcohol by S_N1

(2) secondary alcohol by S_N2

(3) tertiary alcohol by S_n2

(4) secondary alcohol by S_NI.

Set. (1)

Reaction proceeds through carbocation formation as 3° carbocation is highly stable, hence reaction proceeds through $S_N 1$ with 3° alcohol.

- 2. The first ionization potential of Na is 5.1 eV. The value of electron gain enthalpy of Na' will be:
 - (1) = 5.1 eV

(3) -10.2 eV

(3) + 2.55 eV

(4) -2.55 eV

Sed. (1)

 $Na = \frac{AB + +5.5eV}{4} Na^{*} + e^{-}$, here the backward reaction releases same amount of energy and known as

Electron gain enthalpy.

- Stability of the species Li₂, Li₂ and Li₂ increases in the order of:
 - (1) $Li_2 < Li_2 < Li_3$

(3) $Li_2 < Li_2^* < Li_2^*$

C0: $Li_2 < Li_3 < Li_2^*$

(4) $Li_2 < Li_2^* < Li_2^*$

Sol. (1)

 $Li_2(6) = \sigma Is^2 \dot{\sigma} Is^2 \sigma 2s^2$

B.O. =
$$\frac{4-2}{2} = 1$$

 $\text{Li}_{2}^{+}(5) = \sigma \ln^{2} \hat{\sigma} \ln^{2} \sigma 2 s^{1}$

B. O.
$$=\frac{3-2}{2}=0.5$$

 $Li_2^{\alpha}(7) = \sigma i s^2 \dot{\sigma} i s^2 \sigma 2 s^2 \dot{\sigma} 2 s^4$

B.O.
$$=\frac{4-3}{2}=0.5$$

Li2 is more stable than Li2 because Li2 has more numbers of antibonding electrons.

- The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2(M)HCl will be:
 - (I) 1.00 M (3) 0.975 M

(2) 1.75 M (4) 0.875 M

Sol. (4 M

3.

 $M_1V_1 + M_2V_2 = MV$

$$M = \frac{M_1V_1 + M_2V_2}{V} = \frac{0.5 \times 750 + 2 \times 250}{1000}$$

M = 0.875

- Which of the following is the wrong statement?

 (1) O₁ molecule is bent
- (2) Ozone is violet-black in solid state (4) ONCI and ONO are not isoelectronic
- (3) Ozone is diamagnetic gas

 Sol. (All the options are correct statements)
 - (I) Correct, as O O is bent.

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