Name

Date

Pd \_\_\_\_

# Unit 7: Balancing Chemical Reactions- Worksheet 2 Balance the following equations by inserting the proper coefficients.

- \_2\_SO<sub>2</sub> + \_1\_O<sub>2</sub> → \_2\_SO<sub>3</sub> Synthesis
- 2. 2 CH<sub>4</sub> + 3 O<sub>2</sub> → 2 CO + 4 H<sub>2</sub>O Single R
- 2\_P + \_3\_Cl<sub>2</sub> → \_2\_PCl<sub>3</sub> Synthesis
- 4. 2\_CO + 1\_O<sub>2</sub> → 2\_CO<sub>2</sub> Synthesis
- 2\_CH<sub>4</sub> + 1\_O<sub>2</sub> → 2\_CH<sub>3</sub>OH Synthesis
- 2\_Li + \_1\_Br<sub>2</sub> → \_2\_LiBr Synthesis
- 7.  $_{2}Al_{2}O_{3} \rightarrow _{4}Al + _{3}O_{2} Decomp$
- 8. \_2\_Na + \_2\_H<sub>2</sub>O → \_2\_NaOH + \_1\_H<sub>2</sub> Single R
- 9.  $_{6}$  CO<sub>2</sub> +  $_{6}$  H<sub>2</sub>O  $\rightarrow$   $_{1}$  C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> +  $_{6}$  O<sub>2</sub> Rev. Comb
- 10. 1 H2SO4 + 2 NaCl → 2 HCl + 1 Na2SO4 Double R

Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:

- 11. 3 H<sub>2</sub> + 1 SO<sub>2</sub> → 1 H<sub>2</sub>S + 2 H<sub>2</sub>O Double R
- 12.  $_{2}^{2}$ CaCO<sub>3</sub> +  $_{2}^{2}$ SO<sub>2</sub> +  $_{1}^{1}$ O<sub>2</sub>  $\rightarrow$   $_{2}^{2}$ CaSO<sub>4</sub> +  $_{2}^{2}$ CO<sub>2</sub>No
- 13. \_2\_AgNO<sub>3</sub> + \_1\_CaCl<sub>2</sub> → \_2\_AgCl + \_1\_Ca(NO<sub>3</sub>)<sub>2</sub> Double R
- 14. 2 HCl + 1 Ba(OH)<sub>2</sub> → 1 BaCl<sub>2</sub> + 2 H<sub>2</sub>O Double R
- 15. \_1\_H<sub>3</sub>PO<sub>4</sub> + \_3\_NaOH → \_1\_Na<sub>3</sub>PO<sub>4</sub> + \_3\_H<sub>2</sub>O Double R
- 16. \_1\_Pb(NO<sub>3</sub>)<sub>2</sub> + \_2\_KI → \_1\_PbI<sub>2</sub> + \_2\_KNO<sub>3</sub> Double R
- 17. \_3\_CuO + \_2\_NH<sub>3</sub> → \_1\_N<sub>2</sub> + \_3\_Cu + \_3\_H<sub>2</sub>O Single R
- 18.  $1_{C_2H_5OH} + 3_{O_2} \rightarrow 2_{CO_2} + 3_{H_2O} Com$
- 19.  $_{2}C_{2}H_{6} + _{3}O_{2} \rightarrow _{2}CH_{3}COOH + _{2}H_{2}ONO$
- 20.  $_{3}NO_{2} + _{1}H_{2}O \rightarrow _{2}HNO_{3} + _{1}NO NO$

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