

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	159.6

Modeling Chemistry Unit 7 2 Answers

RJ Alexander



Modeling Chemistry Unit 7 2 Answers:

Data Modeling for Metrology and Testing in Measurement Science Franco Pavese, Alistair B. Forbes, 2008-12-16

The aim of this book is to provide firstly an introduction to probability and statistics especially directed to the metrology and testing fields and secondly a comprehensive newer set of modelling methods for data and uncertainty analysis that are generally not considered yet within mainstream methods. The book brings for the first time a coherent account of these newer methods and their computational implementation. They are potentially important because they address problems in application fields where the usual hypotheses that are at the basis of most of the traditional statistical and probabilistic methods, for example relating to normality of the probability distributions, are frequently not fulfilled to such an extent that an accurate treatment of the calibration or test data using standard approaches is not possible. Additionally, the methods can represent alternative ways of data analysis allowing a deeper understanding of complex situations in measurement. The book lends itself as a possible textbook for undergraduate or postgraduate study in an area where existing texts focus mainly on the most common and well known methods that do not encompass modern approaches to calibration and testing problems. The book is structured in such a way to guide readers with only a general interest in measurement issues through a series of review papers from an initial introduction to modelling principles in metrology and testing to the basic principles of probability in metrology and statistical approaches to certainty assessment.

Chemistry in the Community American Chemical Society, 2006-01-31. This laboratory based text centres itself around decision making activities where students apply their chemistry knowledge to realistic situations. This fifth edition includes more photographs, new drawings and new design.

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