

### Part C: Electron Configuration

12. How many electrons can each level hold? 1st = \_\_\_\_\_ 2nd = \_\_\_\_\_ 3rd = \_\_\_\_\_

13. What term is used for the electrons in the outermost shell or energy level? \_\_\_\_\_

14. Scientists use two types of diagrams to show the electron configuration for atoms. Follow your teacher's directions to complete the diagrams.

Sulfur

Atomic # = 16

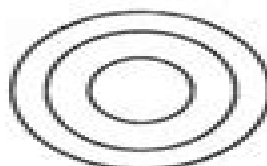
Atomic Mass = 32

Protons = \_\_\_\_\_

Neutrons = \_\_\_\_\_

Electron = \_\_\_\_\_

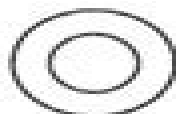
Bohr Diagram  
Shows all electrons



Lewis Structure  
Shows valence electrons

S

15. Calculate the missing information and then draw the Bohr Diagram and Lewis Structure for each element.



Li

Atomic # = 3

Mass # = 7

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Ne

Atomic # = 10

Mass # = 20

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Mg

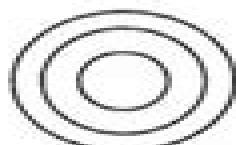
Atomic # = 12

Mass # = 24

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Cl

Atomic # = 17

Mass # = 35

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



He

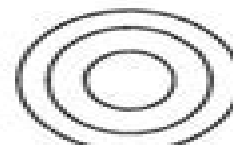
Atomic # = 2

Mass # = 4

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_



Si

Atomic # = 14

Mass # = 28

# of P = \_\_\_\_\_

# of N = \_\_\_\_\_

# of E = \_\_\_\_\_

16. Answer the questions below based on the elements in question #15.

(1) Which elements had a filled outermost shell? \_\_\_\_\_

(2) Which element would be most likely to lose electrons in a chemical bond? \_\_\_\_\_

(3) Which element would be most likely to gain electrons in a chemical bond? \_\_\_\_\_

(4) Which elements are not likely to bond with other elements? \_\_\_\_\_ Why? \_\_\_\_\_

# Ionic And Covalent Bonds Worksheet 7 2 Answers

**R Pring**



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Glencoe Science McGraw-Hill Staff, 2001-08 The Nature of Matter Gr. 5-8, Chemistry, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards

Science Spectrum Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03

**Discover Science: Teacher's annotated edition**, 1991

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