

# Chemistry: Ionic & Covalent Bonds Unit Study Guide

## Covalent Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Covalent Compounds.

- Bondable
- Large / Small
- Atoms / Ions
- High / Low
- Conducts / Does not Conduct Electricity
- Solid / Liquid / Gas at Room Temperature

2. Write the prefix following number.

Number	Prefix
1	
2	
3	
4	
5	

3. What are the two

4. Complete the chart

Formula	Name
$\text{SF}_6$	
	tr
N-L	
	4

5. What does VSEPR

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

## Ionic Bonding & Nomenclature Study Guide

1. Circle the correct word that would complete the statements below regarding Ionic Compounds.

- Bond between a Metal / Nonmetal and a Metal / Nonmetal
- Large / Small Electronegativity Difference
- The (Metal / Nonmetal) form Cations which (Gain / Lose) Electrons
- The (Metal / Nonmetal) form Anions which (Gain / Lose) Electrons
- High / Low Melting & Boiling Point
- Conducts / Does not Conduct Electricity in Solution
- Solid / Liquid / Gas at Room Temperature

2. What charge would the following periodic table families have in a chemical bond?

Family	Charge
alkali metals	
alkali earth	
halogens	
transition	
metals	
nonmetals	

3. Write the roman numeral that is represented by the following numbers.

Roman Numeral	Number
	1
	2
	3
	4
	5
	6

4. Complete the chart below by filling in the missing names.

Cation	Anion	Formula	Name
$\text{Mg}^{2+}$	$\text{Br}^-$		
		$\text{CuO}$	
			iron (II) chloride
		$(\text{NH}_4)_3\text{PO}_4$	

# Ionic Covalent Bonding Test Study Guide Answers

**H Kauffman**



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