

- Use the VSEPR theory to predict molecular geometry of BrF_3 .
 a) tetrahedral c) trigonal bipyramidal e) octahedral
 b) see-saw d) square pyramidal
- Use the VSEPR theory to predict molecular geometry of CO_3^{2-} .
 a) bent c) trigonal pyramidal e) trigonal planar
 b) tetrahedral d) T-shaped
- What is the bond angle between the O-N-O bonds in NO_2^+ ion?
 a) 180° c) 118° e) 90°
 b) 120° d) 90°
- What is the molecular shape of SeO_4^{2-} ?
 a) see-saw c) trigonal planar e) T-shaped
 b) trigonal pyramidal d) tetrahedral
- What is the molecular shape of the SF_6^- ion?
 a) linear c) trigonal planar e) T shaped
 b) bent d) tetrahedral
- For XeF_4 , the valence electron geometry around the central atom is _____.
 a) trigonal planar c) trigonal bipyramidal e) square planar
 b) tetrahedral d) octahedral
- The molecular shape for XeF_4 is _____.
 a) trigonal planar c) trigonal bipyramidal e) square planar
 b) tetrahedral d) octahedral
- The F-Xe-F bond angle in XeF_4 molecule is _____.
 a) 90° c) none of these
 b) 109.5° d) 180°

For the following 4 questions, the following molecular substances are referred to. Any substance may be used more than once in answering the questions. CO_2 , SO_2 , SiF_4 , XeF_4 .

9. The species which has tetrahedral as the valence electron geometry is _____.
a) CO_2 c) SF_4 e) none of these
b) SO_2 d) XeF_4
10. The species which has linear molecular geometry is _____.
a) CO_2 c) SF_4 e) none of these
b) SO_2 d) XeF_4
11. The species which is polar is _____.
a) CO_2 c) SF_4 e) none of these
b) SO_2 d) XeF_4
12. The species which has two lone pairs of valence electrons on the central atom is _____.
a) CO_2 b) SO_2 c) SF_4 d) XeF_4

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Zhe Xu



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