

## Lab 5: RATE AND ORDER OF A CHEMICAL REACTION: IODINATION OF ACETONE

### I. Introduction

The purpose of this experiment is to study a reaction that, in the vicinity of room temperature, proceeds at a moderate, relatively easily measured rate. For a given reaction, the rate typically increases with an increase in the concentration of any reactant. The relation between rate and concentration is a remarkably simple one in many cases. The objectives for this experiment include to determine the orders of reaction with respect to acetone and hydrogen ion; to confirm that the order of reaction with respect to iodine is zero; and to evaluate the rate constant  $k$  for the mixtures prepared at room temperature. For this experiment, we took safety precautions using laboratory materials and solutions, including wearing personal protective equipment.

### II. Results Analysis

#### A. Data

Table 1. Reaction Rate Determination

Mixture Number	Initial Concentrations*			Time for reaction (s)	Temperature (°C)**	Reaction rate (M/s or mole/L · s)
	Acetone (M)	Iodine (M)	HCl (M)			
I (D.1)	5 mL	5 mL	5 mL	71.92	25	
II (D.2)	10 mL	5 mL	5 mL	35.10	25	
III (D.3)	5 mL	5 mL	10 mL	39.11	25	
IV (D.4)	5 mL	2.5 mL	5 mL	42.54	25	

Data Table 1, Group C - Iodination of Acetone: Time to Complete Reaction

		Temperature °C	Time it Took for Mixture Color to Disappear (seconds)
Mixture 1	(D.1)	25	71.92
Mixture 2	(D.2)	25	35.10
Mixture 3	(D.3)	25	39.11
Mixture 4	(D.4)	25	42.54

#### B. Discussion

Based on my results, I observed how the freezing point depression is directly related to the molecular weight of the solute in the solution. Using this relationship, the molecular weight of

# Iodination Of Acetone Report

**Katrin Zwirgmaier**



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