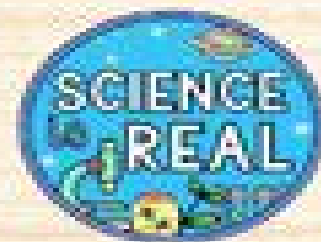


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TYPES OF CHEMICAL BONDS
Atomic Combinations

1. IONIC BONDING: Occurs between a metal and a non-metal. Electrons are transferred from the metal to the non-metal, forming oppositely charged ions that attract each other.

2. COVALENT BONDING: Occurs between two non-metals. Atoms share their outer electrons to form a stable electron shell.



3. METALLIC BONDING: Occurs between metal atoms. The atoms are arranged in a regular lattice, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the lattice together.

ATOMIC STRUCTURE

Atoms consist of a central nucleus containing protons and neutrons, surrounded by electrons in shells. The number of protons equals the number of electrons in a neutral atom.

PERIODIC TABLE

The periodic table is a tabular arrangement of the chemical elements, organized based on their atomic number, electron configuration, and recurring chemical properties.



ATOMIC STRUCTURE

Atoms are made up of three subatomic particles: protons, neutrons, and electrons. Protons and neutrons are located in the nucleus, while electrons orbit the nucleus in shells.

IONIC BONDING




Occurs between a metal and a non-metal. The metal atom loses one or more electrons to form a positive ion, and the non-metal atom gains these electrons to form a negative ion. The resulting ions are held together by strong electrostatic forces.

COVALENT BONDING

Occurs between two non-metals. The atoms share their outer electrons to form a stable electron shell. This sharing of electrons is represented by a pair of dots or a line between the atoms.

METALLIC BONDING

Occurs between metal atoms. The atoms are arranged in a regular lattice, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the lattice together.



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IONIC BONDING

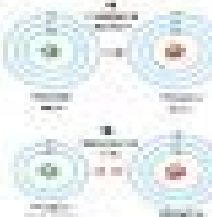
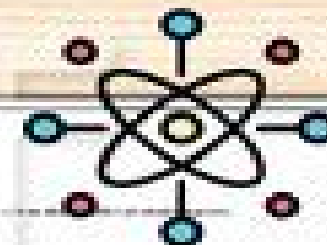
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A Gutmann



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