

## Final Study Guide

**Buoyancy** - ability of a fluid to exert an upward force on object immersed in a fluid. To float, an object must displace an amount of water that has a mass greater than the mass of the object. This usually happens in things that are hollow or filled with air.

**Viscosity** - a fluid's resistance to flow or how thick, or hard to pour a fluid is.

**Density** - mass per unit volume. Density =  $\frac{\text{mass}}{\text{volume}}$

**Physical Property** - a property that can be measured without affecting the substance.

Ex: color, shape, size, mass, anything that can be observed with your senses

**Chemical Property** - a property that tells us how the substance interacts with other substances, the substance is changed in the process.

Ex: Flammability, reactivity

**Physical change** - a change that does not change the identity of the substances involved.

Ex: dissolving, cutting, tearing, melting, freezing

**Chemical change** - A change that does change the identity of the substances involved. Bonds must be broken and new bonds formed.

**Signs of chemical change**: Bubbles (production of a gas), precipitate (production of a new solid), new color, new odor, change in temperature.

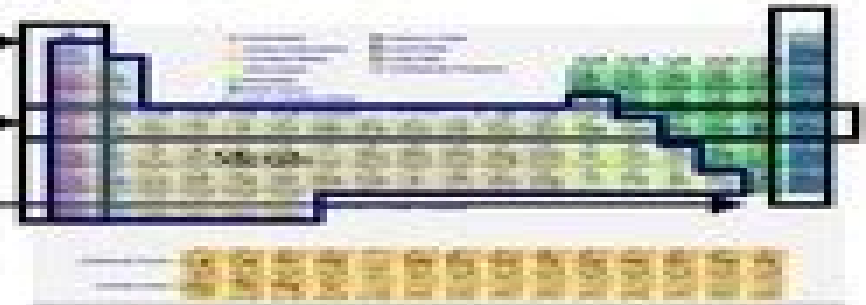
Ex: Combustion, oxidation

**Periodic Table**

**Groups**: the columns of the periodic table, elements in the same group have similar characteristics

**Periods**: the rows of the periodic table, tell you how many orbitals or shells that an atom of a particular element have

**Noble Gases** - unreactive gases in group 18



symbol



Name

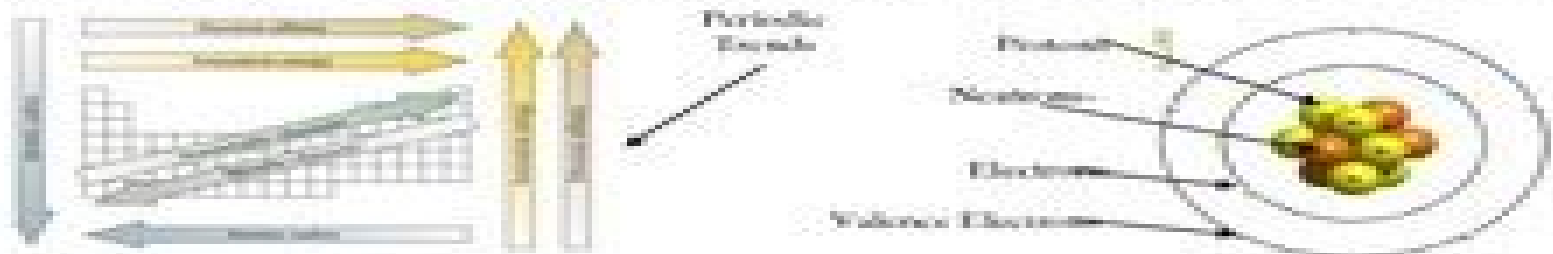
**Atomic number**: equal to the number of protons in atoms of that element

**Atomic mass**: weighted average mass of all the isotopes of an element

**Mass number**: Atomic number rounded to the nearest whole number, number of protons + number of neutrons

**Metals**: The elements left of the jagged line. They are usually good conductors, are malleable, are ductile, and react to magnets. All are solids except mercury.

**Nonmetals**: The elements to the right of the jagged line. Their properties are opposite to those of metals.



**Proton**: Positively charged particle found in the nucleus

**Neutron**: Neutral particle found in the nucleus

**Electrons**: Negatively charged particles found orbiting the nucleus

Glynnis Lundberg

# Ipc Chemistry Review

**O García**



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## **Table of Contents IpC Chemistry Review**

1. Understanding the eBook IpC Chemistry Review
  - The Rise of Digital Reading IpC Chemistry Review
  - Advantages of eBooks Over Traditional Books
2. Identifying IpC Chemistry Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an IpC Chemistry Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from IpC Chemistry Review
  - Personalized Recommendations
  - IpC Chemistry Review User Reviews and Ratings
  - IpC Chemistry Review and Bestseller Lists
5. Accessing IpC Chemistry Review Free and Paid eBooks
  - IpC Chemistry Review Public Domain eBooks

- IpC Chemistry Review eBook Subscription Services
- IpC Chemistry Review Budget-Friendly Options
- 6. Navigating IpC Chemistry Review eBook Formats
  - ePub, PDF, MOBI, and More
  - IpC Chemistry Review Compatibility with Devices
  - IpC Chemistry Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of IpC Chemistry Review
  - Highlighting and Note-Taking IpC Chemistry Review
  - Interactive Elements IpC Chemistry Review
- 8. Staying Engaged with IpC Chemistry Review
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers IpC Chemistry Review
- 9. Balancing eBooks and Physical Books IpC Chemistry Review
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection IpC Chemistry Review
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine IpC Chemistry Review
  - Setting Reading Goals IpC Chemistry Review
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of IpC Chemistry Review
  - Fact-Checking eBook Content of IpC Chemistry Review
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

## 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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