1. (4 points)

The specific heat (s) of iron (Fe) is 0.40 J/(g K). Calculate the heat effect q when 20 g of iron is heated from 100 °C to 200 °C. Include the sign!

 $q = m.s.\Delta T$ for $\Delta T \Delta K = \Delta^{\circ}C$ q = 20x0.40x(200-100) = +800 J

You must include the units and sign in the answer!

2. (4 points)

Reaction: $Mg(s) + \frac{1}{2}O_3(g) \rightarrow MgO(s)$ $\Delta H = -600 \text{ kJ/mol}$

The Molar Mass of Mg = 24.0 g/mol.

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

 $2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol})x(-600)(kJ/\text{mol}) = -60 \text{ kJ}$

Again, your answer must include both the sign and the units.

 (2 points) Which one of the following four reactions is a formation reaction? <u>Circle</u> the correct answer.

a)
$$H_2O(I) + \frac{1}{2}O_2 \longrightarrow H_2O_2(I)$$

b)
$$SO_2(g) + \frac{1}{2}O_2 \rightarrow SO_2(g)$$

c)
$$Cl(g) + H(g) \rightarrow HCl(g)$$

d)
$$S(s) + O_2(g) \rightarrow SO_2(g)$$

Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur, so e.g. Cl_2 not Cl_4 and H_2 not H_3 !

CHEM 101 Section L52 Quiz #6 b

1. (4 points)

The specific heat (s) of copper (Cu) is 0.38 J/(g.K). Calculate the heat effect q when 10 g of copper is heated from 100 °C to 300 °C. Include the sign!

 $q = m.s.\Delta T$ for $\Delta T \Delta K = \Delta^*C$ q = 10x0.38x(300-100) = +760 J

You must include the units and sign in the answer!

2. (4 points)

Reaction: $Ca(s) + \frac{1}{2}O_2(g) \rightarrow CaO(s)$

$$\Delta H = -650 \text{ kJ/mol}$$

The Molar Mass of Ca = 40.0 g/mol.

Calculate the enthalpy change ΔH when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

 $8.0 \text{ g Ca} = 8.0/40/ = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol})x(-650)(kJ/\text{mol}) = -130 \text{ kJ}$

Again, your answer must include both the sign and the units.

 (2 points) Which one of the following reactions is a formation reaction? <u>Circle</u> the correct answer.

a)
$$Br(g) + H(g) \rightarrow HBr(g)$$

b)
$$H_2O(I) + \frac{1}{2}O_2 \rightarrow H_2O_2(I)$$

c)
$$2P(s) + 2\%O_2(g) \rightarrow P_2O_3(s)$$

d)
$$SO_2(g) + \frac{1}{2}O_2 \rightarrow SO_3(g)$$

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RD Boyd

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A Guide to the Literature of Chemistry Evan Jay Crane, Austin McDowell Patterson, 1927

Student's Guide, Chemistry, the Central Science James C. Hill, 1991 An Introduction to the World's Oceans Keith A. Sverdrup, Alyn C. Duxbury, Alison Duxbury, 2003 In An Introduction to the World's Oceans Seventh Edition Keith Sverdrup Alyn Duxbury and Alison Duxbury have blended the most comtemporary information and research with basic principles to bring you and your students an unmatched comprehensive introduction to oceanography You will find a significantly revised Seventh Edition that addresses all the latest findings in oceanography What's special about these authors An Introduction to the World's Oceans Seventh Edition contains balanced and comprehensive coverage that comes from each author having strength in different areas of oceanography Oceanography is an eclectic science that examines physical chemical and biological properties of the world's oceans Alison Duxbury has a background in marine biology Alyn Duxbury has a background in physical oceanography and Keith Sverdrup has a background in marine geology geophysics and how oceanography relates to other areas of science The result A well balanced comprehensive introduction to oceanography McGraw Hill has exclusive videos from Scripps Institution of Oceanography These video clips will be brief one to two minute clips and available on either videotape or on the Digital Content Manager CD ROM There will be a total of about 2 hours and 12 minutes worth of these short clips Clips will be available for each chapter of the text and no other company can offer these Mastering Pharmacy Licensure and Certification David C. Kosegarten, Douglas J. Pisano, F. Randy Vogenberg, 2000 videos Integrating knowledge from the core pharmacy basic disciplines this guide presents over 100 intriguing case studies Each study is followed by ten to 15 questions designed to prepare students for the pharmacy board exam Five of the questions are clinically oriented and focus on the 200 commonly used drugs Student's Guide to Brown and LeMay, Chemistry, the Central Science, 2nd Edition James C. Hill, 1981 **Chemistry** James C. Hill,2003 This book assists students through the text material with chapter overviews learning objectives review of key terms cumulative chapter review guizzes and self tests Included are answers to all Student Guide exercises Chapter summaries are correlated to those in the Instructor's Resource

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