

CHEMISTRY 10TH - Multiple Choice Questions

→ CHEMICAL EQUILIBRIUM

CHAPTER# 01

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01. Which one of the following statements is false about dynamic equilibrium?
- (a) It takes place in a closed container
 - (b) Concentration of reactant and products are not changed
 - (c) Rate of the forward reaction is equal to the rate of the reverse reaction
 - (d) ✓ Equilibrium cannot be disturbed by any external stress
02. Consider the following reaction and indicate which of the following best describes equilibrium constant expression K_c . $4\text{NH}_3 + 5\text{O}_2 \rightleftharpoons 4\text{NO} + 6\text{H}_2\text{O}$
- (a) ✓ $K_c = \frac{[\text{NO}]^4 [\text{H}_2\text{O}]^6}{[\text{NH}_3]^4 [\text{O}_2]^5}$
 - (b) $K_c = \frac{[\text{NH}_3]^4 [\text{O}_2]^5}{[\text{NO}]^4 [\text{H}_2\text{O}]^6}$
 - (c) $K_c = \frac{[\text{NH}_3][\text{O}_2]}{[\text{H}_2\text{O}][\text{NO}]}$
 - (d) $K_c = \frac{[4\text{NO}][6\text{H}_2\text{O}]}{[4\text{NH}_3][5\text{O}_2]}$
03. A reaction which is never goes to completion is known as reversible reaction. Reversible reaction is represented by:
- (a) dotted lines
 - (b) single arrow
 - (c) ✓ double arrow
 - (d) double straight line
04. When the magnitude of K_c is small, indicates:
- (a) ✓ Reaction mixture contain most of the reactant
 - (b) Reaction mixture contain most of the product
 - (c) Reaction mixture contain almost equal amount of reactant and product
 - (d) Reaction goes to completion
05. For which system does the equilibrium constant, K_c has units of concentration?
- (a) ✓ $\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$
 - (b) $\text{N}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)}$
 - (c) $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$
 - (d) $\text{CO}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_2\text{O}_{(l)}$

Multiple Questions On Chemical Equilibrium

Yijin Wang



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Table of Contents Multiple Questions On Chemical Equilibrium

1. Understanding the eBook Multiple Questions On Chemical Equilibrium
 - The Rise of Digital Reading Multiple Questions On Chemical Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Questions On Chemical Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Questions On Chemical Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Questions On Chemical Equilibrium
 - Personalized Recommendations
 - Multiple Questions On Chemical Equilibrium User Reviews and Ratings
 - Multiple Questions On Chemical Equilibrium and Bestseller Lists

5. Accessing Multiple Questions On Chemical Equilibrium Free and Paid eBooks
 - Multiple Questions On Chemical Equilibrium Public Domain eBooks
 - Multiple Questions On Chemical Equilibrium eBook Subscription Services
 - Multiple Questions On Chemical Equilibrium Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Multiple Questions On Chemical Equilibrium Compatibility with Devices
 - Multiple Questions On Chemical Equilibrium Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Questions On Chemical Equilibrium
 - Highlighting and Note-Taking Multiple Questions On Chemical Equilibrium
 - Interactive Elements Multiple Questions On Chemical Equilibrium
8. Staying Engaged with Multiple Questions On Chemical Equilibrium
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Questions On Chemical Equilibrium
9. Balancing eBooks and Physical Books Multiple Questions On Chemical Equilibrium
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Questions On Chemical Equilibrium
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Multiple Questions On Chemical Equilibrium
 - Setting Reading Goals Multiple Questions On Chemical Equilibrium
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Questions On Chemical Equilibrium
 - Fact-Checking eBook Content of Multiple Questions On Chemical Equilibrium
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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