

Element	Q	R	V	X	Y	Z
Atomic Radii(nm)	0.136	0.099	0.125	0.117	0.110	0.157
Formulae of oxide	QO	R ₂ O	V ₂ O ₅	XO ₂	Y ₂ O ₃	Z ₂ O
Melting point(°C)	650	110	660	119	44.2	97.8
Conductivity	Good	Poor	Good	Poor	Poor	Good

- (i) From the table which elements could be:
 I. Magnesium.....(1 mark)
 II. Sulphur.....(1 mark)
- (ii) Write the formula of the chloride of X.(1 mark)
- (iii) Arrange the elements as they appear from left to right in the periodic table
 (2marks)

- (iv) What type of bonding would you expect the chloride of Y to exhibit? (1 mark)

- (v) Write a chemical equation for the reaction of oxide of Y with water. (1 mark)

(b) Use the information in the table below to answer the questions that follow.

Element	Ionization Energy in kJ			
	1 st	2 nd	3 rd	4 th
V	320	580	4900	7200
W	430	4300	6200	8300
X	600	1050	1450	16000
Y	7000	9000	10500	14000
Z	200	420	5200	6600

- (i) Name two elements that can be found in the same group? (1 mark)

- (ii) To which group does element W belong? Explain. (2marks)

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M Woodhall



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