Chapter 6

Chapter 6: Thermochemistry

6.2 - Enthalpy and Calorimetry

Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{products} - H_{reactants}$$

Products of a reaction have a greater enthalpy (ΔH):

- ΔH is positive
- · heat is absorbed by the system (endothermic reaction)

Reactants of a reaction have a greater enthalpy (ΔH):

- · AH is negative
- · heat is released by the system (exothermic reaction)

Calorimetry

Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

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Diane Alford, Jennifer Hill

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