

PHYSICS GRADE 10

CHAPTER ONE LINEAR MOTION

Introduction

Study of motion is divided into two;

1. Kinematics
2. Dynamics

In kinematics forces causing motion are disregarded while dynamics deals with motion of objects and the forces causing them.

I. Displacement

Distance moved by a body in a specified direction is called displacement. It is denoted by letter 's' and has both magnitude and direction. Distance is the movement from one point to another. The SI unit for displacement is the metre (m).

II. Speed

This is the distance covered per unit time.

Speed = distance covered / time taken. Distance is a scalar quantity since it has magnitude only. The SI unit for speed is metres per second (m/s or ms⁻¹)

Average speed = total distance covered / total time taken

Other units for speed used are Km/h.

Examples

1. A body covers a distance of 10m in 4 seconds. It rests for 10 seconds and finally covers a distance of 50m in 60 seconds. Calculate the average speed.

Solution

Total distance covered = 10 + 50 = 100m

Total time taken = 4 + 10 + 6 = 20 seconds

Therefore average speed = 100 / 20 = 5m/s

2. Calculate the distance in metres covered by a body moving with a uniform speed of 180 km/h in 30 seconds.

Solution

Distance covered = speed * time

$$= 180 * 1000 / 60 * 60 = 50 \text{ m/s}$$

$$= 50 * 30$$

$$= 1,500 \text{ m}$$

3. Calculate the time in seconds taken a by body moving with a uniform speed of 360km/h to cover a distance of 3,000 km?

Solution

Speed: 360 km/h = 360 * 1000 / 60 * 60 = 100m/s

Time = distance / speed

$$3000 * 1000 / 100$$

$$= 30,000 \text{ seconds.}$$

III. Velocity

This is the change of displacement per unit time. It is a vector quantity.

Velocity = change in displacement / total time taken

The SI units for velocity are m/s

Examples

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KJ Lindholm-Leary



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Table of Contents Mindset Learn Physics Grade 10 Final Exam 2014

1. Understanding the eBook Mindset Learn Physics Grade 10 Final Exam 2014
 - The Rise of Digital Reading Mindset Learn Physics Grade 10 Final Exam 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Mindset Learn Physics Grade 10 Final Exam 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mindset Learn Physics Grade 10 Final Exam 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mindset Learn Physics Grade 10 Final Exam 2014
 - Personalized Recommendations
 - Mindset Learn Physics Grade 10 Final Exam 2014 User Reviews and Ratings
 - Mindset Learn Physics Grade 10 Final Exam 2014 and Bestseller Lists
5. Accessing Mindset Learn Physics Grade 10 Final Exam 2014 Free and Paid eBooks

- Mindset Learn Physics Grade 10 Final Exam 2014 Public Domain eBooks
 - Mindset Learn Physics Grade 10 Final Exam 2014 eBook Subscription Services
 - Mindset Learn Physics Grade 10 Final Exam 2014 Budget-Friendly Options
6. Navigating Mindset Learn Physics Grade 10 Final Exam 2014 eBook Formats
- ePub, PDF, MOBI, and More
 - Mindset Learn Physics Grade 10 Final Exam 2014 Compatibility with Devices
 - Mindset Learn Physics Grade 10 Final Exam 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Mindset Learn Physics Grade 10 Final Exam 2014
 - Highlighting and Note-Taking Mindset Learn Physics Grade 10 Final Exam 2014
 - Interactive Elements Mindset Learn Physics Grade 10 Final Exam 2014
8. Staying Engaged with Mindset Learn Physics Grade 10 Final Exam 2014
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mindset Learn Physics Grade 10 Final Exam 2014
9. Balancing eBooks and Physical Books Mindset Learn Physics Grade 10 Final Exam 2014
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mindset Learn Physics Grade 10 Final Exam 2014
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mindset Learn Physics Grade 10 Final Exam 2014
- Setting Reading Goals Mindset Learn Physics Grade 10 Final Exam 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mindset Learn Physics Grade 10 Final Exam 2014
- Fact-Checking eBook Content of Mindset Learn Physics Grade 10 Final Exam 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used

the amount of product predicted to be recovered was 8 7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 and H_2SO_4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even closer for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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