

SESSION 12: ENERGY & CHEMICAL CHANGE

Key Concepts

In this session we will focus on summarising what you need to know about:

- Activation energy and activation complex
- Energy profile of a reaction
- Enthalpy (ΔH)
- Endothermic and exothermic reactions

X-planation

ENERGY CHANGES DURING CHEMICAL REACTIONS

- Most reactions do not begin until an amount of energy (activation energy) has been added to the reaction mixture.
- The **activation energy** is often called the 'energy hill' which must be 'overcome' by the addition of this amount of energy before a reaction can take place.
- When activation energy is added to the reactants, a so-called **activated complex** is formed.
- Activated complex is a temporary, unstable, high-energy composition of atoms, which represents a transition state between reactants and the products.
- When the activated complex is formed during a reaction, this complex can lead either to the formation of new bonds, i.e. molecules of the products, or to re-formation of the old bonds, thereby returning to being reactants.
- For the reaction



The formation of an activated complex as a transitional step can be represented as follows:



The peak of the energy hill indicates the energy of the activated complex.

- When an activated complex is formed during a reaction, this complex can lead either to the formation of new bonds, i.e. molecules of the products, or to the re-formation of the old bonds thereby returning to being the reactants. This is **reversibility for the reaction**.

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