

- Directions:**  
 1. An element may be assigned a positive or negative oxidation number of zero.  
 2. An atom with a higher electronegativity is assigned a negative oxidation number.  
 3. The sum of all oxidation numbers for a molecule must be zero.

# Ion Practice Set

1. Write an equation that shows how the element is oxidized or reduced.  
 2. Write the net ionic equation for the reaction.  
 3. Write the net ionic equation for the reaction.

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	CHARGE OF ION	TYPE OF ION
1	Sodium	$\text{Na}^+$	11	10	+1	cation
2	Chlorine	$\text{Cl}^-$	17	18	-1	anion
3	Strontium	$\text{Sr}^{2+}$	38	36	+2	cation
4	Phosphorus	$\text{P}^{3-}$	15	18	-3	anion
5	Calcium	$\text{Ca}^{2+}$	20	18	+2	cation
6	Barium	$\text{Ba}^{2+}$	56	54	+2	cation
7	Hydrogen	$\text{H}^+$	1	0	+1	cation
8	Lithium	$\text{Li}^+$	3	2	+1	cation
9	Magnesium	$\text{Mg}^{2+}$	12	10	+2	cation
10	Sulfur	$\text{S}^{2-}$	16	18	-2	anion
11	Selenium	$\text{Se}^{2-}$	34	36	-2	anion
12	Oxygen	$\text{O}^{2-}$	8	10	-2	anion
13	Aluminum	$\text{Al}^{3+}$	13	10	+3	cation
14	Carbon	$\text{C}^{4-}$	6	10	-4	anion
15	Nitrogen	$\text{N}^{3-}$	7	10	-3	anion
16	Fluorine	$\text{F}^-$	9	10	-1	anion
17	Bromine	$\text{Br}^-$	35	36	-1	anion
18	Iodine	$\text{I}^-$	53	54	-1	anion
19	Francium	$\text{Fr}^+$	87	86	+1	cation

# Ion Practice Set Answer Key

**Henry George Denham**



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