

Final Exam Review Multiple-Choice Questions Combined

1. Use the VSEPR theory to predict molecular geometry of IF_5 .
a) tetrahedral c) trigonal bipyramidal e) octahedral
b) see-saw d) square pyramidal
2. Use the VSEPR theory to predict molecular geometry of CO_3^{2-} .
a) bent c) trigonal pyramidal e) trigonal planar
b) tetrahedral d) T-shaped
3. What is the bond angle between the O-N-O bonds in NO_2^+ ion?
a) 180° b) 120° c) 118° d) 90° e) 88°
4. What is the molecular shape of SeO_2 ?
a) see-saw c) trigonal planar e) T-shaped
b) trigonal pyramidal d) tetrahedral
5. What is the molecular shape of the SF_4 ion?
a) linear c) trigonal planar e) T shaped
b) bent d) tetrahedral
6. For XeF_4 , the valence electron geometry around the central atom is _____.
a) trigonal planar c) trigonal bipyramidal e) square planar
b) tetrahedral d) octahedral
7. The molecular shape for XeF_4 is _____.
a) trigonal planar c) trigonal bipyramidal e) square planar
b) tetrahedral d) octahedral
8. The F-Xe-F bond angle in XeF_4 molecule is _____.
a) 90° c) none of these
b) 109.5° d) 180°

For the following 4 questions, the following molecular substances are referred to. Any substance may be used more than once in answering the questions. CO_2 , SO_2 , SiF_4 , XeF_4 .

9. The species which has tetrahedral as the valence electron geometry is _____.
a) CO_2 c) SiF_4 e) none of these
b) SO_2 d) XeF_4
10. The species which has linear molecular geometry is _____.
a) CO_2 c) SiF_4 e) none of these
b) SO_2 d) XeF_4
11. The species which is polar is _____.
a) CO_2 c) SiF_4 e) none of these
b) SO_2 d) XeF_4
12. The species which has two lone pairs of valence electrons on the central atom is
a) CO_2 b) SO_2 c) SiF_4 d) XeF_4

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