

Name: _____

Sci # _____

COVALENT BOND WEBQUEST

DIRECTIONS: Answer the following by going to the notes website. You can go to Mrs. Slowik's website and use the links or type in the addresses for each section if needed.

Activity #1- Introduction to Covalent Bonding

Open [Chemical Bonding](#). Scroll down to the heading "Covalent Bonding." OR
http://www.visionlearning.com/library/module_viewer.php?mid=55

1. As opposed to _____ bonding in which a complete transfer of electrons occur, _____ bonding occurs when two or more elements _____ electrons. Covalent bonding occurs because the atoms in the compound have a similar tendency for electrons (generally to _____ electrons). This most commonly occurs when two _____ bond together. Because both of the nonmetals will want to _____ electrons, the elements involved will share electrons in an effort to _____ their valence shells.

2. How many valence electrons are in one atom of hydrogen?

3. How many valence electrons does hydrogen need to have a full first shell?

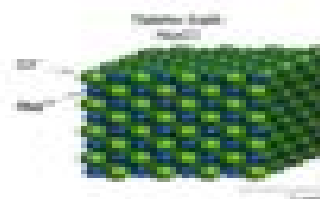
4. How does the hydrogen atom "pick up" another electron?

5. What compound does hydrogen form?

6. How do hydrogen atoms make a covalent bond?

7. Visit the simulation, [Covalent bonding between hydrogen atoms](#). OR
http://www.visionlearning.com/library/flash_viewer.php?fid=1346&mid=55
Describe or draw what you see.

8. Label the substances below as "ionic" or "covalent".



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LM Reder



Ionic And Covalent Bonding Webquest Answer Key:

Ionic and Covalent Bonding James M. Postma, Julian Roberts, J. Leland Hollenberg, 2000-01-15 *The Answer Key*
Rachel Turoscy, 2018-08-09 The Answer Key A Comprehensive Explanation of Problem Solving Methods for General Chemistry Success Volume 1 is a concise and accessible textbook that covers the critical information a student needs to understand the basic mathematics used in chemistry courses The book provides easy to understand step by step instructions for solving general chemistry problems The book begins with chapters dedicated to problem solving methodology and unit conversions In subseq

The Covalent Bond in Crystals. Ii. Partially Ionic Binding J. C. Phillips, CHICAGO UNIV ILL., 1967 The general theory of covalent bonding developed in AD 657 432 is extended to heteropolar crystals In diatomic crystals the theory depends on two parameters an energy gap for each atom or an energy gap for the medium and a polarizability parameter ϵ The latter can be determined by comparison with empirical pseudopotential form factors the value obtained for GaAs and ZnSe which brings the form factors of the latter into agreement with the ion core form factors of Animalu and Heine is $\epsilon = 0.05$ or 0.05 This polarizability parameter can be used to estimate the effective charge e^* defined by Callen which describes the splitting of longitudinal and transverse optic modes of vibration at $q = 0$ The result in GaAs is $e^* = 0.14$ or 0.04 in rough agreement with the experimental value of 0.20 or 0.03 The experimental values for e^* may be used to estimate the gap parameters for each atom when complete data are available Author

Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H bonding Noncovalent bonding Polar Polar covalent bond Polar molecules Covalent bonds Polarity of covalent bonds Polar covalent bond Polarity of molecules Noncovalent bonding Interaction Noncovalent bonding Van der Waals Forces Covalent bonds Subdivision of covalent bonds Covalent Bond Classification Polar covalent bond Hybrids Noncovalent bonding Electrostatic Interactions Polar covalent bond Nonpolar

molecules Covalent Bond Classification Other uses Covalent bonds History Noncovalent bonding Dipole Dipole Noncovalent bonding Protein Folding Structure and much more **Atoms, Electrons, Structure and Bonding** Primrose

Kitten, 2017-06-19 I'm constantly telling you the best way to learn is by practicing questions so I've made you a book full of practice questions Multiple choice questions to reflect the style of exam questions activities to complete equations for you to balance compounds for you to work out the formula for lots of things that you need to recall and practice long answer exam style questions This book is not designed as a text book or revision guide but as a workbook There are lots of good and bad expensive and free revision guides out there on my YouTube channel and other great websites So there is no point in me adding to the masses All the teaching all the new content is available for free on my YouTube channel this book is for you to practice and learn The best way to approach this is to watch the teaching video and make notes or after class try a section and check the answers Any corrections that are needed after the book is published will be listed on my website www.primrosekitten.com these will be corrected in the next version of the book Answers are provided for the sections where you need to work out the answers for yourself not the sections where you are just filling things in from a video or website

Atoms Electrons Structure and Bonding Workbook Topics Covered are Some of this content has also been published in the Summer Start to A Level Chemistry and a practice exam paper

Atomic Structure 20 Multiple choice questions Properties of Ionic Compounds 15 multiple choice questions Reference table of common ions formulae Formula of Ionic Compounds 65 formulas to work out Drawing Ionic Bonding 10 Compounds Simple Covalent Bonding 20 multiple choice questions Drawing Covalent Bonding 10 Compounds Summary Table for the 4 Different Types of Bonding Electron configurations Drawing electron configurations Drawing Electron Configurations Spot the mistake Electronic Configuration 20 multiple choice questions Exceptions to the Octet Rule Oxidation Numbers 20 multiple choice questions Balancing Equations using the oxidation numbers method 20 to practice Salt Equations 20 equations to complete and balance Shapes of Molecules Investigation Shapes of Molecules and Bond Angles 20 multiple choice questions Electronegativity and Bond Polarity Investigation Intermolecular Bonding 10 multiple choice questions Electrons Structure and Bonding Practice Exam Questions Answers

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