

### Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Model 1 (on the next page) to complete the following table.

	Metal or Nonmetal?	Is the number of protons the same in the atom and the ion?	Is the number of neutrons the same in the atom and the ion?	Is the number of electrons the same in the atom and the ion?	Charge on the ion
Lithium	Metal	Yes	Yes	No	1+
Magnesium	Metal	Yes	Yes	No	2+
Aluminum	Metal	Yes	Yes	No	3+
Fluorine	Non-Metal	Yes	Yes	No	1-
Oxygen	nonmetal	Yes	Yes	No	2-
Nitrogen	non-metal	Yes	Yes	No	3-

2. Based on the table, what distinguishes a neutral atom from an ion?  
a neutral atom has the same amount of  $P^+$  and  $E^-$
3. Examine the isotope symbols in Model 1.
- a) Where is the ion charge located in the isotope symbol?  
the charge of an ion comes after the charge number
- b) Is a charge indicated on the neutral atoms? If yes, where is it located?  
A charge isn't indicated in neutral atoms b/c they don't have a charge and  $P^+ - E^- = 0$ .
4. Which subatomic particle carries a positive charge?  
 $E^+$ ,  $Mg$ ,  $Al$
5. Which subatomic particle carries a negative charge?  
 $F$ ,  $O$ ,  $N$
6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using a positive ion example and a negative ion example from Model 1.

$$\begin{aligned} \text{Charge of an ion} &= \text{number of protons} - \text{number of electrons} \\ \text{Nitrogen} &= P^+ - E^- \\ &= 7 - 10 \\ &= -3 (3-) \end{aligned}$$

$$\begin{aligned} \text{Aluminum} &= P^+ - E^- \\ &= 13 - 10 \\ &= +3 (3+) \end{aligned}$$

# Ions Pogil Activity Answers

**M Mark**



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## **Table of Contents Ions Pogil Activity Answers**

1. Understanding the eBook Ions Pogil Activity Answers
  - The Rise of Digital Reading Ions Pogil Activity Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ions Pogil Activity Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ions Pogil Activity Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ions Pogil Activity Answers
  - Personalized Recommendations
  - Ions Pogil Activity Answers User Reviews and Ratings
  - Ions Pogil Activity Answers and Bestseller Lists
5. Accessing Ions Pogil Activity Answers Free and Paid eBooks
  - Ions Pogil Activity Answers Public Domain eBooks
  - Ions Pogil Activity Answers eBook Subscription Services
  - Ions Pogil Activity Answers Budget-Friendly Options

6. Navigating Ions Pogil Activity Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ions Pogil Activity Answers Compatibility with Devices
  - Ions Pogil Activity Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ions Pogil Activity Answers
  - Highlighting and Note-Taking Ions Pogil Activity Answers
  - Interactive Elements Ions Pogil Activity Answers
8. Staying Engaged with Ions Pogil Activity Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ions Pogil Activity Answers
9. Balancing eBooks and Physical Books Ions Pogil Activity Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ions Pogil Activity Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ions Pogil Activity Answers
  - Setting Reading Goals Ions Pogil Activity Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ions Pogil Activity Answers
  - Fact-Checking eBook Content of Ions Pogil Activity Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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