

**123****III**

Total No. of Questions – 21

Regd.

Total No. of Printed Pages – 2

No.

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Part – III
CHEMISTRY, Paper – I
(English Version)

Time : 3 Hours]

[Max. Marks : 60

Note : Read the following instructions carefully :

- (1) Answer all the questions of Section – ‘A’. Answer any six questions in Section – ‘B’ and any two questions in Section – ‘C’.
- (2) In Section – ‘A’, questions from Sr. Nos. 1 to 10 are of “Very short answer type”. Each question carries two marks. Every answer may be limited to two or three sentences. Answer all these questions at one place in the same order.
- (3) In Section – ‘B’, questions from Sr. Nos. 11 to 18 are of “Short answer type”. Each question carries four marks. Every answer may be limited to 75 words.
- (4) In Section – ‘C’, questions from Sr. Nos. 19 to 21 are of “Long answer type”. Each question carries eight marks. Every answer may be limited to 300 words.
- (5) Draw labelled diagrams, wherever necessary, for questions in Section – ‘B’ and Section – ‘C’.

SECTION – A**10 × 2 = 20****Note :** Answer all questions :

1. Why KO_2 paramagnetic?
2. Describe the importance of plaster of paris.
3. H_2O has a higher boiling point than that of H_2S . Give reasons.
4. State First Law of thermodynamics.
5. State Graham’s law of diffusion.
6. Calculate the weight of 0.1 mole of Sodium carbonate.

[1 of 2]

123

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 - Advantages of eBooks Over Traditional Books
2. Identifying Ipe Chemistry Previous Question Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ipe Chemistry Previous Question Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ipe Chemistry Previous Question Papers
 - Personalized Recommendations

- Ipe Chemistry Previous Question Papers User Reviews and Ratings
- Ipe Chemistry Previous Question Papers and Bestseller Lists
- 5. Accessing Ipe Chemistry Previous Question Papers Free and Paid eBooks
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 - Ipe Chemistry Previous Question Papers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ipe Chemistry Previous Question Papers
 - Highlighting and Note-Taking Ipe Chemistry Previous Question Papers
 - Interactive Elements Ipe Chemistry Previous Question Papers
- 8. Staying Engaged with Ipe Chemistry Previous Question Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ipe Chemistry Previous Question Papers
- 9. Balancing eBooks and Physical Books Ipe Chemistry Previous Question Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ipe Chemistry Previous Question Papers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ipe Chemistry Previous Question Papers
 - Setting Reading Goals Ipe Chemistry Previous Question Papers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ipe Chemistry Previous Question Papers
 - Fact-Checking eBook Content of Ipe Chemistry Previous Question Papers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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