



$$\frac{.13g \text{ Mg} | 1 \text{ mol}}{24.3g} = .00535 \text{ mol Mg}$$

5. Determine the number of moles of chlorine that reacted.

$$\frac{.33g \text{ Cl} | 1 \text{ mol}}{35.5g} = .00930 \text{ mol Cl}$$

6. Determine the simplest (empirical) ratio of moles that reacted.

$$\frac{.00930 \text{ mol Cl}}{.00535 \text{ mol Mg}} = 1.74 \quad \text{MgCl}_x$$



$$.00535 \text{ mol Mg}$$



# Lab Determining An Empirical Formula Mgcl Answers

**Tao Wei**



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