

## Additivity of Heats of Reaction: Hess's Law

### Advanced Study Questions:

Name \_\_\_\_\_

One of the reactions below is the same as the combination (addition) of the other two reactions. Therefore, according to Hess's Law, the heat of reaction,  $\Delta H$ , of that reaction should be equal to the sum of the heats of reaction for the other two.

- (1)  $\text{NaOH(s)} \rightarrow \text{Na}^{\text{+}}(\text{aq}) + \text{OH}^{-}(\text{aq})$   $\Delta H_1$
- (2)  $\text{NaOH(s)} + \text{H}^{\text{+}}(\text{aq}) + \text{Cl}^{-}(\text{aq}) \rightarrow \text{H}_2\text{O(l)} + \text{Na}^{\text{+}}(\text{aq}) + \text{Cl}^{-}(\text{aq})$   $\Delta H_2$
- (3)  $\text{Na}^{\text{+}}(\text{aq}) + \text{OH}^{-}(\text{aq}) + \text{H}^{\text{+}}(\text{aq}) + \text{Cl}^{-}(\text{aq}) \rightarrow \text{H}_2\text{O(l)} + \text{Na}^{\text{+}}(\text{aq}) + \text{Cl}^{-}(\text{aq})$   $\Delta H_3$

Determine which of the three reactions above may result from the direct addition of the other two. Clearly show how two equations add up to give the third one. (*Use equations as they are: do not reverse, adjust coefficients, etc.*)

For the reaction obtained as a sum of two others (called the overall reaction), show how can its  $\Delta H$  be expressed using  $\Delta H$  for two other reactions? *You do not need values, use symbols only.*

In this experiment, sodium hydroxide is delivered in two ways, as solid or in an aqueous solution. Which sample contains **larger number of moles** of NaOH: 2.65 g solid NaOH or 55.2 mL of 1.15M NaOH solution? *Show full work including units.*

## Lab 35 Heats Of Reaction Answers

**Theodore Rolly Ball**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the Theodore Rolly Ball text.

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