

Name : \_\_\_\_\_ Date : \_\_\_\_\_



## pH and pOH Worksheet



1. A solution has a hydrogen ion concentration of  $2.8 \times 10^{-6}$  M. What is the pOH of its solution ?

2. A solution has a  $[\text{OH}^-]$  of  $5.8 \times 10^{-7}$ . What is the pH of this solution ?

3. A 450 mL beaker is 0.00045 M HCl. What is the pH of this solution ?

4. A 320 mL solution contains 2.30 mg of NaOH. What is the pH of this solution ?

5. Complete the following table.

| pH   | $[\text{H}^+]$       | $[\text{OH}^-]$      | pOH |
|------|----------------------|----------------------|-----|
| 3.5  |                      |                      |     |
|      | $5.8 \times 10^{-7}$ |                      |     |
|      |                      | $4.2 \times 10^{-2}$ |     |
|      |                      |                      | 8.2 |
|      | $4.2 \times 10^{-5}$ |                      |     |
|      |                      |                      | 2.4 |
| 10.1 |                      |                      |     |
|      |                      | $7.2 \times 10^{-3}$ |     |

# Mcgraw Hill Ph Poh Worksheet

**Sasha Hu**



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