

Name : \_\_\_\_\_ Date : \_\_\_\_\_



## pH and pOH Worksheet



1. A solution has a hydrogen ion concentration of  $2.8 \times 10^{-6}$  M. What is the pOH of its solution ?

2. A solution has a  $[\text{OH}^-]$  of  $5.8 \times 10^{-7}$ . What is the pH of this solution ?

3. A 450 mL beaker is 0.00045 M HCl. What is the pH of this solution ?

4. A 320 mL solution contains 2.30 mg of NaOH. What is the pH of this solution ?

5. Complete the following table.

pH	$[\text{H}^+]$	$[\text{OH}^-]$	pOH
3.5			
	$5.8 \times 10^{-7}$		
		$4.2 \times 10^{-2}$	
			8.2
	$4.2 \times 10^{-5}$		
			2.4
10.1			
		$7.2 \times 10^{-3}$	

# Mcgraw Hill Ph Poh Worksheet

**N Noddings**



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