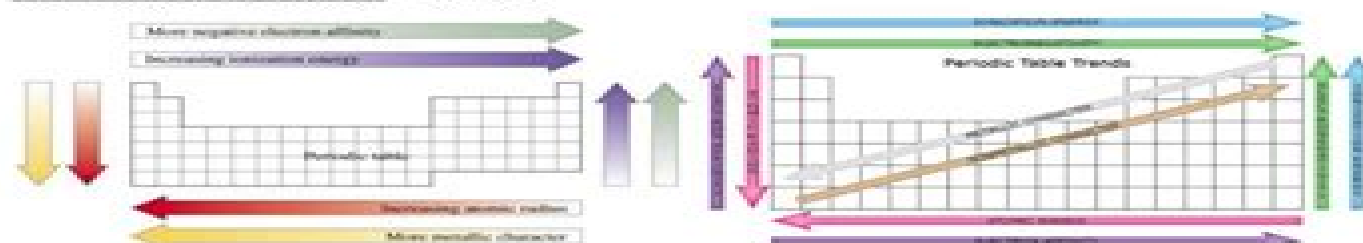


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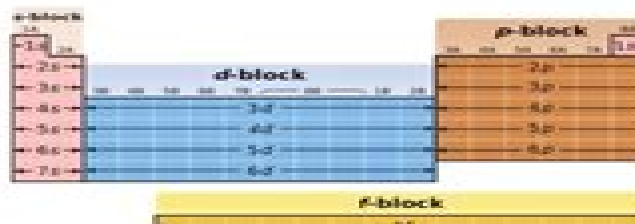
### The Nature of Matter - Trends



### Atomic Radius:

- Moving down  $\rightarrow$  size increases
  - Why? Because the # of energy shells increases and a new inner energy level is added between the nucleus and the atom's outer electron cloud
    - With each additionally inner level, the outer cloud is more shielded from the attractive force of the nucleus
    - *Electron Cloud:* a model of the atom that states we cannot know where electrons are at all times, but they are more likely to remain in certain areas (cloud) of the atom; think, vague
    - *What Affects EC Size?:* Good question! An increase in # of energy levels increases the probability that the outer electrons spend time away from the nucleus, and so the size increases
- Moving from left to right  $\rightarrow$  size decreases
  - Why? Because  $Z_{eff}$  increases and the force of attraction between the nucleus and the valence electrons is stronger (thus, the size decreases)
    - *What is  $Z_{eff}$ ?* the effective nuclear charge:  $Z$  (# of protons (atomic #)) -  $S$  (# of core electrons (shielding electrons)); conceptually, the amount of charge "felt" or "seen" by outer valence electrons
- Neutral atoms tend to get bigger as you go down a group and decrease as you go across a period
- When a neutral atom gains an electron (forming an anion), the atomic radius increases; When a neutral atom loses an electron (forming a cation), the atomic radius decreases

### Ionic Radius:



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