

FAR EASTERN UNIVERSITY - MAKATI
MANAGEMENT SERVICES - INTEGRATED REVIEW
ACADEMIC YEAR 2025 - 2026 | 1st Semester
REVIEWER

1. The basic rule of the time value of money is:
 - A. investments will always be worth more tomorrow than they are today
 - B. it's always wiser to save a dollar for tomorrow than to spend it today
 - C. **a dollar in hand today is worth more than a dollar promised at some time in the future**
 - D. all of the above express an aspect of the basic rule of time value of money
2. The present value of a future amount:
 - A. will always be less than the future amount
 - B. can be calculated precisely if the discount rate and number of periods is known
 - C. is worth less than the future value
 - D. **both a. and b. above are true**
3. The process of determining the present value of a future cash flow is known as:
 - A. Compounding
 - B. **Discounting**
 - C. Annuitizing
 - D. Budgeting
4. Net working capital is defined as
 - A. a ratio measure of liquidity best used in cross-sectional analysis.
 - B. the portion of the firm's assets financed with short-term funds.
 - C. current liabilities minus current assets.
 - D. **current assets minus current liabilities.**
5. In working capital management, risk is measured by the probability that a firm will become
 - A. liquid.
 - B. **technically insolvent.**
 - C. unable to meet long-term obligations.
 - D. less profitable.
6. If a firm uses an aggressive financing strategy,
 - A. **it increases return and increases risk.**
 - B. it increases return and decreases risk.
 - C. it decreases return and increases risk
 - D. it decreases return and decreases risk.

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what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
a helpful way to remember this

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oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + \text{I}^-$ 2 $\text{O}_5 + 5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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state and oxidation number concepts to identify redox reactions and to keep track of electrons

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reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} oxidized reducing agent O_2 to O^{2-} reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+1} oxidized red ag
 S^0 to S^{2-} reduced ox ag $3\text{Cs} + \text{Br}_2$