

Chemistry Answer key

Carbon and Its Compounds

A. Short Answer Conceptual Questions:

1. Covalent bonds: Formed by sharing electrons between atoms to achieve a stable octet (or duplet for hydrogen). Example: Methane (CH_4) - Carbon shares 4 valence electrons with 4 hydrogen atoms, each forming a single covalent bond. (A simple dot and cross diagram would be very helpful here).
2. Catenation: The ability of carbon atoms to form long chains or rings by bonding with other carbon atoms. This property is due to carbon's small size and high electronegativity. Examples: Ethane (C_2H_6), Benzene (C_6H_6), and long hydrocarbon chains in polymers like polyethylene.
3. Saturated hydrocarbons: Contain only single bonds between carbon atoms. They are less reactive. Example: Ethane (C_2H_6). Unsaturated hydrocarbons: Contain double or triple bonds between carbon atoms. They are more reactive due to the presence of pi bonds. Examples: Ethene (C_2H_4) (double bond), Ethyne (C_2H_2) (triple bond).
4. Homologous series: A series of organic compounds with the same general formula, similar chemical properties, and successive members differing by a CH_2 group. Alkanes (general formula $\text{C}_n\text{H}_{2n+2}$): Methane (CH_4), Ethane (C_2H_6), Propane (C_3H_8), Butane (C_4H_{10}).
5. Isomerism: The phenomenon where compounds have the same molecular formula but different structural arrangements of atoms. Example: Hexane (C_6H_{14}) has five isomers: n-hexane, 2-methylpentane, 3-methylpentane, 2,2-dimethylbutane, and 2,3-dimethylbutane. (Drawing the structures is essential for a complete answer).

B. Long Answer Application-Based Questions:

6. Methane (CH_4): Carbon (4 valence electrons) shares 1 electron with each of the 4 hydrogen atoms (1 valence electron), forming 4 single covalent bonds. (Diagram essential). Ethyne (C_2H_2): Each carbon (4 valence electrons) shares 3 electrons with the other carbon and 1 electron with a hydrogen atom, forming a triple bond between the carbons and a single bond with each hydrogen. (Diagram essential).
7. Soaps: Sodium or potassium salts of long-chain fatty acids. They are effective in soft water because they form lather readily. In hard water, they react with calcium and magnesium ions to form insoluble precipitates (scum). Detergents: Ammonium or sulfonate salts of long-chain carboxylic acids. They are effective in hard water because they do not form insoluble precipitates with calcium and magnesium ions.

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