

# Sulfur clock lab

## Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	Reaction setup
B	0.5 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.5 mL, 0.01 M $\text{KSCN}$
C	0.25 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.25 mL, 0.01 M $\text{KSCN}$
D	0.125 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.125 mL, 0.01 M $\text{KSCN}$

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the calculated amount of  $\text{H}_2\text{O}_2$  solution to the reaction mixture using a graduated cylinder.
6. Add exactly 0.01 M  $\text{KSCN}$  solution to the reaction mixture.
7. Prepare and use the standard solutions.
8. Prepare solutions A, B, C, D and E as shown:
  - a. Transfer 0.1 mL of 0.1 M  $\text{H}_2\text{O}_2$  solution to the reaction mixture.
  - b. Add 0.01 M  $\text{KSCN}$  solution to the reaction mixture using a graduated cylinder.
  - c. Dilute to the 10 mL mark using distilled water.
  - d. Use the flask and record the initial concentration.
9. Transfer solution A into the graduated cylinder as 0.1 mL flask. Use the same flask and add any additional water.

# Lab Sulfur Clock Answers

**Lori Wilde**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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