

Experiment No. 03

Date:

Name of the experiment: Standardization of HCl Solution by NaOH Solution.

Principle

An acid-base titration involves the addition of a **titrant** solution to an **analyte** solution. In titrimetry, chemicals used as reference solutions known as primary standards or secondary standards.

A **primary standard** is a reagent that is extremely pure, stable, has no water of hydration, and has a high molecular weight. A **secondary standard** solution refers to a solution that has its concentration measured by titration with a primary standard solution.

The secondary standard chemicals are not absolute pure and may contain some impurities that come during their synthesis. Therefore the secondary standard chemical should standardize by using the primary pure chemicals.

In the present experiment, NaOH solution is to be used secondary standard chemical and needs to be standardized by the primary standard chemicals. Here the primary standard chemical is Oxalic acid. After being standardized by the Oxalic acid, NaOH is to be used for the standardization of HCl.

The end point of the solution will be determined by an indicator, phenolphthalein.

Reaction:**Apparatus:**

1. Electrical balance
2. Spatula
3. Volumetric flask
4. Measuring cylinder
5. Burette
6. Pipette and pipette filler
7. Conical flask
8. Beaker
9. Funnel

Chemicals / Reagents:

1. 0.1 N Sodium hydroxide (NaOH) solution
2. 0.1 N Oxalic acid (dihydrate) [$\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$]
3. 0.1 N Hydro-chloric acid (HCl) solution
4. 0.5% Phenolphthalein indicator
5. Distilled Water

Calculation (Sample): Full calculations required to make the reagents, for the experiment must be written down:

Gram equivalent weight of $\text{H}_2\text{C}_2\text{O}_4 = (126.08 \div 2) \text{ gm} = 63.040 \text{ gm}$

100 mL 0.1N $\text{H}_2\text{C}_2\text{O}_4 = (63.040 \times 100 \times 0.1) \div 1000 \text{ gm} = Y \text{ gm}$ (for 100% pure $\text{H}_2\text{C}_2\text{O}_4$)

Reagents and their preparations:

1. **0.1 N 100 mL Oxalic acid solution:** Completely dissolve 0.636 gm of Oxalic acid (99%) in a 100 mL volumetric flask with a small volume of DW. Finally adjust the volume up to 100 mL mark by adding distilled water Q.S.
2. **0.1 N 1000 mL NaOH solution:** Completely dissolve 4.12 gm of NaOH (97%) in a 1000 mL volumetric flask with a small volume of DW. Finally adjust the volume up to 1000 mL mark by adding distilled water Q.S.
3. **0.1N 100 mL HCl solution:** Take 1.14 mL of HCl (32%) and completely dissolve it in 100 mL volumetric flask in 100 mL DW.
4. **0.5% 100 mL Phenolphthalein solution:** 50 mL Ethanol+ 0.5 g Phenolphthalein+ 50 mL DW.

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Table of Contents Laboratory Procedure Manual Format

1. Understanding the eBook Laboratory Procedure Manual Format
 - The Rise of Digital Reading Laboratory Procedure Manual Format
 - Advantages of eBooks Over Traditional Books
2. Identifying Laboratory Procedure Manual Format
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Laboratory Procedure Manual Format
 - User-Friendly Interface
4. Exploring eBook Recommendations from Laboratory Procedure Manual Format
 - Personalized Recommendations
 - Laboratory Procedure Manual Format User Reviews and Ratings
 - Laboratory Procedure Manual Format and Bestseller Lists
5. Accessing Laboratory Procedure Manual Format Free and Paid eBooks
 - Laboratory Procedure Manual Format Public Domain eBooks
 - Laboratory Procedure Manual Format eBook Subscription Services
 - Laboratory Procedure Manual Format Budget-Friendly Options
6. Navigating Laboratory Procedure Manual Format eBook Formats
 - ePub, PDF, MOBI, and More
 - Laboratory Procedure Manual Format Compatibility with Devices
 - Laboratory Procedure Manual Format Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Laboratory Procedure Manual Format
 - Highlighting and Note-Taking Laboratory Procedure Manual Format
 - Interactive Elements Laboratory Procedure Manual Format
8. Staying Engaged with Laboratory Procedure Manual Format

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Laboratory Procedure Manual Format
- 9. Balancing eBooks and Physical Books Laboratory Procedure Manual Format
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Laboratory Procedure Manual Format
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Laboratory Procedure Manual Format
 - Setting Reading Goals Laboratory Procedure Manual Format
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Laboratory Procedure Manual Format
 - Fact-Checking eBook Content of Laboratory Procedure Manual Format
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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