

SL Paper 2

This question is about thermal properties of matter.

a. Explain, in terms of the energy of its molecules, why the temperature of a pure substance does not change during melting. [3]

b. Three ice cubes at a temperature of 0°C are dropped into a container of water at a temperature of 22°C . The mass of each ice cube is 25 g and the mass of the water is 330 g. The ice melts, so that the temperature of the water decreases. The thermal capacity of the container is negligible. [4]

The following data are available.

$$\text{Specific latent heat of fusion of ice} = 3.3 \times 10^5 \text{ J kg}^{-1}$$

$$\text{Specific heat capacity of water} = 4.2 \times 10^3 \text{ J kg}^{-1} \text{ K}^{-1}$$

Calculate the final temperature of the water when all of the ice has melted. Assume that no thermal energy is exchanged between the water and the surroundings.

Markscheme

a. energy supplied/bonds broken/heat absorbed;

increases potential energy;

no change in kinetic energy (so no change in temperature);

b. use of $M \times 4.2 \times 10^3 \times \Delta\theta$

$$mI = 75 \times 10^{-3} \times 3.3 \times 10^5 / 24750 \text{ J};$$

recognition that melted ice warms and water cools to common final temperature;

3.4°C ;

Examiners report

a. [N/A]

b. [N/A]

Part 2 Simple harmonic oscillations

A longitudinal wave travels through a medium from left to right.

Graph 1 shows the variation with time t of the displacement x of a particle P in the medium.

Graph 1

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Hongru Du



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