#### Colligative Properties



 As a solute is added to a solvent, what happens to the freezing point and the boiling point of the solution?

A) The freezing point decreases and the boiling point increases.

- B) The freezing point increases and the boiling point increases.
- The freezing point increases and the boiling. point decreases.
- D) The freezing point decreases and the boiling. point decreases.
- Compared to the freezing point of 1.0 M KCl(aq) at standard pressure, the freezing point of 1.0 M. CaCla(aq) at standard pressure is

B) higher

Which aqueous solution of KI freezes at the lowest реширетилире?

A) 1 mol of KI in 500. g of water

B) 2 mol of KI in 500. g of water (%)

C) I mol of KI in 1000, g of water

D) 2 mol of KI in 1000, g of water

 A I kilogram sample of water will have the highest freezing point when it contains

A) 1 × 10<sup>21</sup> dissolved particles

(B) 1 × 1017 dissolved particles

C) 1 × 1019 dissolved particles

D) 1 × 10<sup>23</sup> dissolved particles laucs.

 A solution consists of 0.50 mole of CaCl: dissolved in 100, grams of H2O at 25°C. Compared to the boiling point and freezing point of 100. grams of H2O at standard pressure, the solution at standard pressure has

 A) a higher boiling point and a lower freezing. point

B) a lower boiling point and a lower freezing

C) a higher boiling point and a higher freezing.

 D) a lower boiling point and a higher freezing. point

6. Which solution will freeze at the lowest bisimplerature?

A) I mole of sugar in 500 g of water

B) I mole of sugar in 1,000 g of water

C) 2 moles of sugar in 1,000 g of water 2 moles of sugar in 500 g of water

7. Which solution has the highest boiling point at standard pressure?

(A))0.10 M K;PO4(30)

B) 0.10 M KCl(ac)

C) 0.10 M KNO<sub>3</sub>(a<sub>0</sub>)

D) 0.10 M K2SO4(aq)

 Which concentration of a solution of CH1OH in water has the lowest freezing point?

A) 0.001 M

C) 0.01 M

9. How do the boiling point and freezing point of a solution of water and calcium chloride at standard pressure compare to the boiling point and freezing

point of water at standard pressure?

 A) Both the freezing point and boiling point of the solution are higher.

B) The freezing point of the solution is higher and the boiling point of the solution is lower.

C) The freezing point of the solution is lower and the boiling point of the solution is higher.

 D) Both the freezing point and boiling point of the solution are lower.

10. Compared to a 2.0 M aqueous solution of NaCl at I atmosphere, a 3.0 M aqueous solution of NaCl at I atmosphere has a

A) lower boiling point and a higher freezing.

B) higher boiling point and a higher freezing.

C) lower boiling point and a lower freezing provinct.

higher boiling point and a lower freezing point

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Scott C. Dulebohn

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