

THE MOLE: LAB ACTIVITY

Station 1:

A) Determine the mass of one drop of water by measuring the mass of 50 drops of water.

Mass of dish (or beaker)	=	_____	g
Mass of dish (or beaker) + 50 drops water	=	_____	g
Mass of 50 drops water	=	_____	g
Mass of 1 drop of water	=	_____	g

B) How many molecules are there in one drop of water?

Remember 1 mole of a substance is 6.02×10^{23} particles (atoms or molecules).
Therefore, 6.02×10^{23} molecules of water will weigh 18.02 g/mol

- Find # of moles of water for 1 drop $n = (\text{mass})/(\text{Molar Mass})$
- # of molecules in one drop of water $= n_{\text{water}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$
 $= ? \text{ molecules}$

Station 2

Calculate the number of copper atoms in a penny.

- Find # of moles of Cu $n = (\text{mass})/(\text{Molar Mass})$
- # of Cu atoms in a penny $= n_{\text{Cu}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$
 $= ? \text{ atoms}$

Station 3

A. Calculate the number of moles in one spoon of sugar. Assume that sugar is made up entirely of sucrose molecules ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$).

$$M_{\text{sucrose}} = 342.34 \text{ g/mol}$$

$$n_{\text{sugar}} = \frac{M_{\text{sucrose}}}{M_{\text{sucrose}}}$$

$$= \underline{\hspace{2cm}} \text{ mol}$$

B. Calculate the number of molecules of sugar in one spoon of sugar.

$$\# \text{ of molecules sucrose} = n_{\text{sucrose}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}}$$

$$= \underline{\hspace{2cm}} \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}}$$

Mole Lab Chemistry I Acc Answers

Katrin Zwirgmaier



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Berkeley Lab Research Review, 1997 Cooperative Chemistry Lab Manual Cooper, 2005-02 The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations *The Success Manual for General Chemistry* Elizabeth Kean, Catherine Middlecamp, 1986 **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Study Guide for Hein and Arena's Foundations of College Chemistry* Peter Scott, Rachael Enriquez Porter, 1999-04-18 How to Survive (and Even Excel In) General Chemistry Elizabeth Kean, Catherine Middlecamp, 1994 A different kind of book about chemistry which teaches readers the process of learning chemistry not the topic itself Proving a valuable supplement to any introductory text this guide offers inside information to help make chemistry less stressful even enjoyable Includes exercises and sections for self assessment **Chemistry, Student Study Guide** John A. Olmsted, Gregory M. Williams, 2005-02-02 100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors Student Study Guide to accompany Foundations of College Chemistry, 11th Edition Morris Hein, Susan Arena, 2003-09-03 Work more effectively and gauge your progress along the way This Student Study Guide that accompanies Hein's Foundations of College Chemistry 11th Edition contains a review and self evaluation section challenge problems and answers and solutions for all

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 - The Rise of Digital Reading Mole Lab Chemistry I Acc Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mole Lab Chemistry I Acc Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mole Lab Chemistry I Acc Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mole Lab Chemistry I Acc Answers
 - Personalized Recommendations
 - Mole Lab Chemistry I Acc Answers User Reviews and Ratings
 - Mole Lab Chemistry I Acc Answers and Bestseller Lists
5. Accessing Mole Lab Chemistry I Acc Answers Free and Paid eBooks
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 - Interactive Elements Mole Lab Chemistry I Acc Answers
8. Staying Engaged with Mole Lab Chemistry I Acc Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mole Lab Chemistry I Acc Answers
9. Balancing eBooks and Physical Books Mole Lab Chemistry I Acc Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mole Lab Chemistry I Acc Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mole Lab Chemistry I Acc Answers
 - Fact-Checking eBook Content of Mole Lab Chemistry I Acc Answers
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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