

10Lab | Experiment 5: The Mole, Counting by Weighing

Last Name _____ First _____

____ / ____ pts

Postlab Questions

Part 4. Calculations. Converting from Mass to Moles for some common objects.

Using the mass of the iron cube, the 50.0mL of water and the packet of sugar, determine the number of moles of metal, water and sugar. Round of to correct number of significant figures and use scientific notation when appropriate.

J The number of moles of water in 50.0 ml of water

	Part4 J
J1) 50.0 mL of water: Write the mass of the 50-ml water from part J.	1) _____ g 50.0 mL water
J2) Write the molar mass of water.	2) _____ g/mole for water
J3) Calculate the number of moles of water in the 50-mL volume of water. Show your calculations below.	3) _____ moles of water
J4) Calculate the number of water molecules in 50.0 mL. Show your calculations below.	4) _____ molecules of water
J5) Calculate the number of Oxygen and H atoms in 50 mL of water. Show your calculations below.	5a) _____ O atoms in water
	5b) _____ H atoms in water

J The number of moles of sucrose in a packet of sugar

J6) Sucrose (sugar), $C_{12}H_{22}O_{11}$: Write the mass of sucrose.	6) _____ g mass of sucrose.
J7) Write the molar mass of sucrose	7) _____ g/mol for sucrose.
J8) Calculate the number of moles of the sucrose. Show your calculation below.	8) _____ moles of sucrose
J9) Calculate the number of molecules of sucrose in the packet of sugar. Show your calculations below.	9) _____ molecules of sugar
J10) Calculate the number of atoms of C, H, and O in the packet of sucrose. Show your calculations below.	10a) _____ C atoms
	10b) _____ H atoms
	10c) _____ O atoms

Mole Lab Counting And Weighing Answers

Steven S. Zumdahl



Mole Lab Counting And Weighing Answers:

Chemistry Essentials For Dummies John T. Moore, 2019-04-16 Chemistry Essentials For Dummies 9781119591146 was previously published as Chemistry Essentials For Dummies 9780470618363 While this version features a new Dummies cover and design the content is the same as the prior release and should not be considered a new or updated product Whether studying chemistry as part of a degree requirement or as part of a core curriculum students will find Chemistry Essentials For Dummies to be an invaluable quick reference guide to the fundamentals of this often challenging course Chemistry Essentials For Dummies contains content focused on key topics only with discrete explanations of critical concepts taught in a typical two semester high school chemistry class or a college level Chemistry I course from bonds and reactions to acids bases and the mole This guide is also a perfect reference for parents who need to review critical chemistry concepts as they help high school students with homework assignments as well as for adult learners headed back into the classroom who just need to a refresher of the core concepts The Essentials For Dummies Series Dummies is proud to present our new series The Essentials For Dummies Now students who are prepping for exams preparing to study new material or who just need a refresher can have a concise easy to understand review guide that covers an entire course by concentrating solely on the most important concepts From algebra and chemistry to grammar and Spanish our expert authors focus on the skills students most need to succeed in a subject

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images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo

Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms

Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet

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