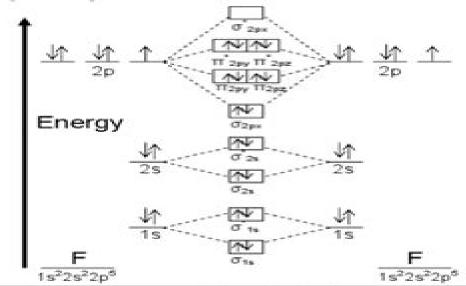
These problems are for practice in drawing your molecular orbital diagrams, molecular electron configurations and determining bond order.

The following questions pertain to the F2 molecule:

 A) Draw the molecular orbital energy diagram for this molecule. Label all of the orbitals specifically.



B) Give the molecular electron configuration for the molecule:

$$(\sigma_{1s})^2 (\sigma_{1s})^2 (\sigma_{2s})^2 (\sigma_{2s})^2 (\sigma_{2px})^2 (\pi_{2p})^4 (\pi_{2p})^4$$

C) Determine the bond order for the molecule:

Bond Order = 
$$(10 - 8)/2 = 1$$

D) Indicate whether the species is paramagnetic or diamagnetic:

Diamagnetic: No unpaired electrons

F) Compare the relative stability of this molecule to F2\* and F2:

# **Molecular Orbital Practice Answers**

**Princeton Review (Firm)** 

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