

Chemistry – Unit 6 Worksheet 1

We have observed evidence that atoms of different elements are different, the most noticeable being their predictable characteristic chemical behaviour. Atoms of one element cannot be transformed into atoms of another element, which demonstrates that each element has a characteristic chemical behaviour. However, when elements react, combinations of atoms are formed by the rearrangement of atoms. This process is called a chemical reaction. We will now examine the patterns that occur in the nature of atoms which show that a substance is either an element or a compound of two or more elements.

1. Write the formulae and draw the particle diagrams for each compound. The ratio of atoms in each compound is given.

| Atomic structure formula | 16 hydrogen 16 oxygen | 12 hydrogen 12 oxygen | 2 aluminium 3 chlorine | 1 hydrogen 1 chlorine |
|--------------------------|--------------------------|--------------------------|---------------------------|--------------------------|
| | H_2O | H_2O_2 | Al_2Cl_3 | HCl |

particle



diagram

| Atomic structure formula | 12 carbon 16 oxygen | 1 hydrogen 1 oxygen | 2 carbon 1 oxygen | 2 carbon 1 oxygen |
|--------------------------|-------------------------------------|------------------------|------------------------|------------------------|
| | $\text{C}_6\text{H}_{12}\text{O}_6$ | H_2O | C_2H_4 | C_2H_2 |

particle



diagram

| Atomic structure formula | 12 hydrogen 16 oxygen | 12 hydrogen 16 oxygen | 1 hydrogen 1 chlorine | 1 hydrogen 1 chlorine |
|--------------------------|--------------------------|--------------------------|--------------------------|--------------------------|
| | H_2O_2 | H_2O | HCl | HCl |

particle



diagram

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Namedate Modeling Chemistry Unit 6 Review Answers

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Glencoe Mathematics ,2001 **Integrated Mathematics** Holt McDougal,1998

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Table of Contents Namedate Modeling Chemistry Unit 6 Review Answers

1. Understanding the eBook Namedate Modeling Chemistry Unit 6 Review Answers
 - The Rise of Digital Reading Namedate Modeling Chemistry Unit 6 Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Namedate Modeling Chemistry Unit 6 Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Namedate Modeling Chemistry Unit 6 Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Namedate Modeling Chemistry Unit 6 Review Answers
 - Personalized Recommendations
 - Namedate Modeling Chemistry Unit 6 Review Answers User Reviews and Ratings
 - Namedate Modeling Chemistry Unit 6 Review Answers and Bestseller Lists

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 - Namedate Modeling Chemistry Unit 6 Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Namedate Modeling Chemistry Unit 6 Review Answers
 - Highlighting and Note-Taking Namedate Modeling Chemistry Unit 6 Review Answers
 - Interactive Elements Namedate Modeling Chemistry Unit 6 Review Answers
8. Staying Engaged with Namedate Modeling Chemistry Unit 6 Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Namedate Modeling Chemistry Unit 6 Review Answers
9. Balancing eBooks and Physical Books Namedate Modeling Chemistry Unit 6 Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Namedate Modeling Chemistry Unit 6 Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Namedate Modeling Chemistry Unit 6 Review Answers
 - Setting Reading Goals Namedate Modeling Chemistry Unit 6 Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Namedate Modeling Chemistry Unit 6 Review Answers
 - Fact-Checking eBook Content of Namedate Modeling Chemistry Unit 6 Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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