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IGCSE A\* MODEL ANSWER

0620 CHEMISTRY

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Paper 42



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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	kJ mol <sup>-1</sup>	kJ mol <sup>-1</sup>
bond energy in kJ/mol	440	435

Use the bond energies in Table 3.1 and  $\Delta H$  to calculate the bond energy of one C-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.
 
$$\begin{aligned}
 & (50H) + 3(H-H) \\
 & 440 + 3(440)
 \end{aligned}
 \quad \underline{1660} \quad \text{kJ}$$
- Calculate the energy released when bonds form in the products.
 
$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -90 \\
 & 1660 - x = -90 \\
 & x = 1660 + 90
 \end{aligned}
 \quad \underline{1750} \quad \text{kJ}$$
- Calculate the energy of one C-H bond.
 
$$\begin{aligned}
 & 6 \times \text{C-H} = 1750 \\
 & \text{C-H} = \frac{1750}{6}
 \end{aligned}
 \quad \underline{292} \quad \text{kJ/mol}$$

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