

# Limiting Reactant and Percentage Yield

- 1 Calcium hydroxide reacts with hydrochloric acid according to the following reaction:



- a Suppose 6.3 mol of HCl reacts with 2.8 mol  $\text{Ca(OH)}_2$ . Which substance is the limiting reactant?
- b How many moles of the excess reactant remain?

- 2 Aluminum oxidizes according to the following unbalanced reaction: (First, balance it)



- a 0.048 mol of powdered Al is placed in a container containing 0.030 mol of  $\text{O}_2$ . What is the limiting reactant?
- b How many moles of excess reactant remain?

- 3 In the reaction between CO and  $\text{Fe}_3\text{O}_4$ , the theoretical yield in an experiment is calculated to be 47.2 g Fe. When a careless chemistry student experiments, the actual yield is 42.9 g Fe. Calculate the percentage yield.

# Limiting Reactant And Percent Yield Lab Answers

**LM Reder**



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