

Name Key

1. Hydrogen, H_2 , is used as rocket fuel. The hydrogen is burned in oxygen, O_2 , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen, H_2 , is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2$$

$$\frac{99.2 \text{ mol } H_2}{2 \text{ mol } H_2} \times \frac{-484 \text{ kJ}}{1} = -24,006 \text{ kJ}$$

$-24,006 \text{ kJ}$
 ~~$-24,006 \text{ kJ}$~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is ΔH for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$. How much heat must be added to 5.2 g of iron at 25°C to raise its temperature to 225°C ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left(\frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine, N_2H_4 , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate ΔH for reaction (1) given the data below.



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1. Understanding the eBook Kp Chemistry Practice Problems
 - The Rise of Digital Reading Kp Chemistry Practice Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Kp Chemistry Practice Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kp Chemistry Practice Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kp Chemistry Practice Problems
 - Personalized Recommendations
 - Kp Chemistry Practice Problems User Reviews and Ratings
 - Kp Chemistry Practice Problems and Bestseller Lists
5. Accessing Kp Chemistry Practice Problems Free and Paid eBooks
 - Kp Chemistry Practice Problems Public Domain eBooks
 - Kp Chemistry Practice Problems eBook Subscription Services
 - Kp Chemistry Practice Problems Budget-Friendly Options

6. Navigating Kp Chemistry Practice Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Kp Chemistry Practice Problems Compatibility with Devices
 - Kp Chemistry Practice Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kp Chemistry Practice Problems
 - Highlighting and Note-Taking Kp Chemistry Practice Problems
 - Interactive Elements Kp Chemistry Practice Problems
8. Staying Engaged with Kp Chemistry Practice Problems
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kp Chemistry Practice Problems
9. Balancing eBooks and Physical Books Kp Chemistry Practice Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kp Chemistry Practice Problems
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Kp Chemistry Practice Problems
 - Setting Reading Goals Kp Chemistry Practice Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Kp Chemistry Practice Problems
 - Fact-Checking eBook Content of Kp Chemistry Practice Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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