

Stoichiometry

● Limiting Reactant & Percentage Yield ●

Answers

Stoichiometry (Page 8) Level 5: Limiting Reactants

- 5) How many grams of NaCl can be produced from the reaction when 50 g solution of HCl reacts with 48 grams of NaOH? The chemical equation is



Mole of NaOH	Mole of HCl	Mole of NaCl	Mole of H ₂ O
$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$
$n = \frac{50}{40} = 1.25 \text{ mol}$	$n = \frac{48}{36.5} = 1.318 \text{ mol}$	$n = \frac{m}{58.5}$	$n = \frac{m}{18}$
$n = 1.25 \text{ mol}$	$n = 1.318 \text{ mol}$	$n = 2.5 \text{ mol}$	$n = 1.25 \text{ mol}$

- 6) How many liters of CO₂ at STP can be produced when 10 g of benzene (C₆H₆) reacts with 50 g of O₂? The chemical equation is unbalanced.



Mole of C ₆ H ₆	Mole of O ₂	Mole of CO ₂	Mole of H ₂ O
$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$
$n = \frac{10}{78} = 0.128 \text{ mol}$	$n = \frac{50}{32} = 1.5625 \text{ mol}$	$n = \frac{m}{44}$	$n = \frac{m}{18}$
$n = 0.128 \text{ mol}$	$n = 1.5625 \text{ mol}$	$n = 0.768 \text{ mol}$	$n = 0.384 \text{ mol}$

Answers

Stoichiometry (Page 12) Level 6: Percentage Yield

1. If 15.8 grams of methane (CH₄) undergo complete combustion in excess oxygen to produce 29.4 grams of water, what is the percentage yield of this reaction? The chemical equation is balanced.



Mole of CH ₄	Mole of O ₂	Mole of CO ₂	Mole of H ₂ O
$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$
$n = \frac{15.8}{16} = 0.9875 \text{ mol}$	$n = \frac{29.4}{18} = 1.633 \text{ mol}$	$n = \frac{m}{44}$	$n = \frac{29.4}{18} = 1.633 \text{ mol}$
$n = 0.9875 \text{ mol}$	$n = 1.633 \text{ mol}$	$n = 0.9875 \text{ mol}$	$n = 1.975 \text{ mol}$

Percentage (%) Yield

$$\% = \frac{\text{Actual}}{\text{Theoretical}} \times 100 \quad \% = \frac{29.4}{35.2} \times 100 = 83.5\%$$

Answer: **83.5%**

2. If 78.3 grams of sodium carbonate (Na₂CO₃) reacts with 833 mL of 1.12 M solution of hydrochloric acid to produce 49.7 grams of sodium chloride, what is the limiting reactant and the percentage yield of this reaction? The chemical equation is unbalanced.



Mole of Na ₂ CO ₃	Mole of HCl	Mole of NaCl	Mole of CO ₂	Mole of H ₂ O
$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$	$n = \frac{m}{M}$
$n = \frac{78.3}{106} = 0.7387 \text{ mol}$	$n = 0.833 \text{ L} \times 1.12 \text{ mol/L} = 0.93196 \text{ mol}$	$n = \frac{49.7}{58.5} = 0.8496 \text{ mol}$	$n = \frac{m}{44}$	$n = \frac{m}{18}$
$n = 0.7387 \text{ mol}$	$n = 0.93196 \text{ mol}$	$n = 0.8496 \text{ mol}$	$n = 0.477 \text{ mol}$	$n = 0.477 \text{ mol}$

Limiting Reactant: **Hydrochloric acid**

Answer Key available

Limiting Reactants And Percentage Yield Quiz

**Burns A. Ralph, Charles Baker, Claire
Baker**



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Table of Contents Limiting Reactants And Percentage Yield Quiz

1. Understanding the eBook Limiting Reactants And Percentage Yield Quiz
 - The Rise of Digital Reading Limiting Reactants And Percentage Yield Quiz
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants And Percentage Yield Quiz
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants And Percentage Yield Quiz
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants And Percentage Yield Quiz
 - Personalized Recommendations
 - Limiting Reactants And Percentage Yield Quiz User Reviews and Ratings
 - Limiting Reactants And Percentage Yield Quiz and Bestseller Lists
5. Accessing Limiting Reactants And Percentage Yield Quiz Free and Paid eBooks
 - Limiting Reactants And Percentage Yield Quiz Public Domain eBooks
 - Limiting Reactants And Percentage Yield Quiz eBook Subscription Services
 - Limiting Reactants And Percentage Yield Quiz Budget-Friendly Options
6. Navigating Limiting Reactants And Percentage Yield Quiz eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactants And Percentage Yield Quiz Compatibility with Devices
 - Limiting Reactants And Percentage Yield Quiz Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants And Percentage Yield Quiz
 - Highlighting and Note-Taking Limiting Reactants And Percentage Yield Quiz
 - Interactive Elements Limiting Reactants And Percentage Yield Quiz
8. Staying Engaged with Limiting Reactants And Percentage Yield Quiz

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Limiting Reactants And Percentage Yield Quiz
- 9. Balancing eBooks and Physical Books Limiting Reactants And Percentage Yield Quiz
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants And Percentage Yield Quiz
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reactants And Percentage Yield Quiz
 - Setting Reading Goals Limiting Reactants And Percentage Yield Quiz
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reactants And Percentage Yield Quiz
 - Fact-Checking eBook Content of Limiting Reactants And Percentage Yield Quiz
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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