

Limiting Reactant, Theoretical Yield, Actual Yield, & Percent Yield

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24



- If we have 53.2 g of Na and 65.8 g of Cl_2 , what are the limiting reactant and theoretical yield?

$$53.2\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} \times \frac{2\text{ mol NaCl}}{2\text{ mol Na}} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 135.23$$

$$65.8\text{ g Cl}_2 \times \frac{1\text{ mol Cl}_2}{70.9\text{ g Cl}_2} \times \frac{2\text{ mol NaCl}}{1\text{ mol Cl}_2} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 110.5$$

Limiting Reagent And Percent Yield 9 3 Review

ML Morrison



Limiting Reagent And Percent Yield 9 3 Review:

Limiting Reagent And Percent Yield 9 3 Review Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the energy of words has been evident than ever. They have the ability to inspire, provoke, and ignite change. Such may be the essence of the book **Limiting Reagent And Percent Yield 9 3 Review**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall impact on readers.

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